

Marking Scheme:

(i) Each question is allotted 4 (four) marks for each correct response.

(ii) $\frac{1}{4}$ (one fourth) marks will be deducted for indicating incorrect response of each question. No deduction from the total score will be made if no response is indicated for an item in the answer sheet.

Q.1 Which of the following molecules has trigonal planar geometry

- (1) IF_3 (2) PCl_3
(3) NH_3 (4) BF_3

Q.2 Experiment shows that H_2O has a dipole moment while CO_2 has not. Point out the structures which best illustrate these facts

- (1) $\text{O}=\text{C}=\text{O}$; 
(2) $\text{O}=\text{C}=\text{O}$; $\text{H}-\text{O}-\text{H}$
(3) ; $\text{H}-\text{H}-\text{O}$
(4) ; 

Q.3 The two types of bonds present in B_2H_6 are covalent and

- (1) Three centre bond (2) Hydrogen bond
(3) Two centre bond (4) None of the above

Q.4 What is the correct mode of hybridisation of the central atom in the following compounds :



- (1) sp^2 , sp^3 , d^2sp^3 (2) sp^3 , sp^3d^2 , sp^3d^2
(3) sp , sp^3d , sp^3d^2 (4) sp , sp^2 , sp^3

Q.5 Which one of the following is a correct set with respect to molecule, hybridisation and shape –

- (1) BeCl_2 , sp^2 , linear
(2) BeCl_2 , sp^2 , triangular planar
(3) BCl_3 , sp^2 , triangular planar
(4) BCl_3 , sp^3 , tetrahedral

Q.6 Many ionic crystals dissolve in water because

- (1) Water is an amphiprotic solvent.
(2) Water is a high boiling liquid.
(3) The process is accompanied by a positive heat of solution.

(4) Water decreases the interionic attraction in the crystal lattice due to solvation.

Q.7 The high melting point and insolubility in organic solvents of sulphanic acid are due to its _____ structure.

- (1) Simple ionic (2) Bipolar ionic
(3) Cubic (4) Hexagonal

Q.8 Which compound will have electrovalent bonding

- (1) Ammonia (2) Water
(3) Calcium chloride (4) Chloromethane

Q.9 Which of the following compounds does not follow the octet rule for electron distribution

- (1) PCl_5 (2) PCl_3
(3) H_2O (4) PH_3

Q.10 Sulphuric acid provides a example of

- (1) Co-ordinate bonds
(2) Non-covalent compound
(3) Covalent and co-ordinate bond
(4) Non-covalent ion

Q.11 In the following which one have zero dipole moment

- (1) BF_3 (2) C
(3) BeCl_2 (4) All of these

Q.12 Which molecule has the largest dipole moment

- (1) HCl (2) HI
(3) HBr (4) HF

Q.13 Which of the following is the most polar

- (1) BF_3 (2) CHCl_3
(3) CH_3OH (4) CH_3Cl

Q.14 Compound formed by sp^3d hybridisation will have structure

- (1) Planar (2) Pyramidal
(3) Angular (4) Trigonal bipyramidal

Q.15 Which of the following pairs are isostructural ?

- (1) CH_3^- and CH_3^+ (2) NH_4^+ and NH_3
(3) SO_4^{2-} and BF_4^- (4) NH_2^- and BeF_2

Q.16 In which of the following species is the underlined carbon having sp^3 -hybridisation?

- (1) $\text{CH}_3\text{-}\underline{\text{C}}\text{OOH}$ (2) $\text{CH}_3\text{-}\underline{\text{C}}\text{H}_2\text{NH}_2$
(3) $\text{CH}_3\text{-}\underline{\text{C}}\text{OCH}_3$ (4) $\text{CH}_2\text{=}\underline{\text{C}}\text{-CH}_3$

Q.17 Which of the following statements is true?

- (1) HF is less polar than HBr .
(2) Water does not contain any ions.

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M Woodhall



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