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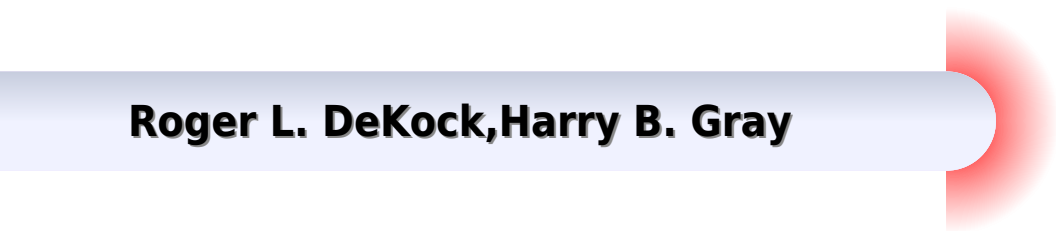
Bonding & Molecular Structure**LEWIS STRUCTURES**

Indicate the total # of VALENCE electrons for each molecule (VE's). Calculate the number of bonded electrons from the skeleton (BE's). Calculate the Remaining Electrons (RE's = VE's - BE's) and distribute the remaining electrons onto the skeletal structure. Finally, satisfy the octets by creating double and triple bonds. Your final structures should be given below and your scratch paper should clearly show your work.

F_2 VE's = <u>14</u> BE's = <u>2</u> RE's = <u>12</u> $\text{:}\ddot{\text{F}}-\ddot{\text{F}}\text{:}$ Shape: <u>linear</u>	H_2 VE's = <u>2</u> BE's = <u>2</u> RE's = <u>0</u> $\text{H}-\text{H}$ Shape: <u>linear</u>	HF VE's = <u>8</u> BE's = <u>2</u> RE's = <u>6</u> $\text{H}-\ddot{\text{F}}\text{:}$ Shape: <u>linear</u>	NH_3 VE's = <u>8</u> BE's = <u>6</u> RE's = <u>2</u> $\begin{array}{c} \text{H} \\ \\ \text{H}-\text{N}-\text{H} \\ \\ \cdot \end{array}$ Shape: <u>trigonal pyramidal</u>
CH_4 VE's = <u>8</u> BE's = <u>8</u> RE's = <u>0</u> $\begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{H} \\ \\ \text{H} \end{array}$ Shape: <u>tetrahedral</u>	C_2H_6 VE's = <u>14</u> BE's = <u>14</u> RE's = <u>0</u> $\begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{H} \\ \quad \\ \text{H} \quad \text{H} \end{array}$ Shape: <u>tetrahedral (both)</u>	AlH_3 VE's = <u>6</u> BE's = <u>6</u> RE's = <u>0</u> $\begin{array}{c} \text{H}-\text{Al}-\text{H} \\ \\ \text{H} \end{array}$ Shape: <u>trigonal planar</u>	BH_3 VE's = <u>6</u> BE's = <u>6</u> RE's = <u>0</u> $\begin{array}{c} \text{H} \\ \\ \text{H}-\text{B}-\text{H} \end{array}$ Shape: <u>trigonal planar</u>
C_2H_4 VE's = <u>12</u> BE's = <u>12</u> RE's = <u>0</u> $\begin{array}{c} \text{H} \quad \text{H} \\ \diagdown \quad \diagup \\ \text{C}=\text{C} \\ \diagup \quad \diagdown \\ \text{H} \quad \text{H} \end{array}$ Shape around one C: <u>trigonal planar</u>	C_2F_4 VE's = <u>36</u> BE's = <u>12</u> RE's = <u>24</u> $\begin{array}{c} \text{:}\ddot{\text{F}} \quad \text{:}\ddot{\text{F}} \\ \diagdown \quad \diagup \\ \text{C}=\text{C} \\ \diagup \quad \diagdown \\ \text{:}\ddot{\text{F}} \quad \text{:}\ddot{\text{F}} \end{array}$ Shape around one C: <u>trigonal planar</u>	CO VE's = <u>10</u> BE's = <u>6</u> RE's = <u>4</u> $\text{:}\ddot{\text{C}}\equiv\ddot{\text{O}}\text{:}$ Shape: <u>linear</u>	O_2 VE's = <u>12</u> BE's = <u>4</u> RE's = <u>8</u> $\text{:}\ddot{\text{O}}=\ddot{\text{O}}\text{:}$ Shape: <u>linear</u>

South Pasadena Bonding And Molecular Structure Worksheet Answers

Roger L. DeKock, Harry B. Gray



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Chemical Structure and Bonding Roger L. DeKock, Harry B. Gray, 1980 *A Pictorial Approach to Molecular Bonding* John G. Verkade, 2012-12-06 With the development of accurate molecular calculations in recent years useful predictions of molecular electronic properties are currently being made. It is therefore becoming increasingly important for the non-theoretically oriented chemist to appreciate the underlying principles governing molecular orbital formation and to distinguish them from the quantitative details associated with particular molecules. It seems highly desirable then that the non-theoretician be able to deduce results of general validity without esoteric mathematics. In this context pictorial reasoning is particularly useful. Such an approach is virtually indispensable if bonding concepts are to be taught to chemistry students early in their careers. Undergraduate chemistry majors typically find it difficult to formulate molecular orbital schemes especially delocalized ones for molecules more complicated than diatomics. The major reason for this regrettable situation is the general impracticability of teaching group theory before students take organic and inorganic courses wherein the applications of these concepts are most beneficial. Consequently many students graduate with the misconception that the ground rules governing bonding in molecules such as NH_3 are somehow different from those which apply to aromatic systems such as C_6H_6 . Conversely seniors and many graduate students are usually only vaguely if at all aware that sigma bonding like extended pi bonding can profitably be described in a delocalized manner when discussing the UV photoelectron spectrum of CH_4 for example. Bonding and Molecular Structure Russell S. Drago, 1971 *Chemical Bonds* Harry B. Gray, 1994-12-05 This profusely illustrated book by a world renowned chemist and award winning chemistry teacher provides science students with an introduction to atomic and molecular structure and bonding. This is a reprint of a book first published by Benjamin Cummings 1973. *Structure of Molecules and the Chemical Bond* Y. K. Syrkin, M. E. Dyatkina, 1950

The Nature of the Chemical Bond and the Structure of Molecules and Crystals Linus Pauling, 1963 The Determination of Molecular Structure P. J. Wheatley, 1953 **The Determination of Molecular Structure** Peter Jaffrey Wheatley, 1962 The Determination of Molecular Structure P. J. Wheatley, 1965 **Structure of Molecules and the Chemical Bond** J. K. Syrkin, Y. K. Syrkin, M. E. Dyatkina, 1964 **Molecular Structure and Energetics; Volume 1: Chemical Bonding Models**, 1986 **MOLECULAR STRUCTURE AND BONDING; THE QUALITATIVE MOLECULAR ORBITAL APPROACH**. B. M. Gimarc, 1979 **Molecular Spectra and Molecular Structure**, 1950

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Table of Contents South Pasadena Bonding And Molecular Structure Worksheet Answers

1. Understanding the eBook South Pasadena Bonding And Molecular Structure Worksheet Answers
 - The Rise of Digital Reading South Pasadena Bonding And Molecular Structure Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying South Pasadena Bonding And Molecular Structure Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an South Pasadena Bonding And Molecular Structure Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from South Pasadena Bonding And Molecular Structure Worksheet Answers
 - Personalized Recommendations
 - South Pasadena Bonding And Molecular Structure Worksheet Answers User Reviews and Ratings
 - South Pasadena Bonding And Molecular Structure Worksheet Answers and Bestseller Lists

5. Accessing South Pasadena Bonding And Molecular Structure Worksheet Answers Free and Paid eBooks
 - South Pasadena Bonding And Molecular Structure Worksheet Answers Public Domain eBooks
 - South Pasadena Bonding And Molecular Structure Worksheet Answers eBook Subscription Services
 - South Pasadena Bonding And Molecular Structure Worksheet Answers Budget-Friendly Options
6. Navigating South Pasadena Bonding And Molecular Structure Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - South Pasadena Bonding And Molecular Structure Worksheet Answers Compatibility with Devices
 - South Pasadena Bonding And Molecular Structure Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of South Pasadena Bonding And Molecular Structure Worksheet Answers
 - Highlighting and Note-Taking South Pasadena Bonding And Molecular Structure Worksheet Answers
 - Interactive Elements South Pasadena Bonding And Molecular Structure Worksheet Answers
8. Staying Engaged with South Pasadena Bonding And Molecular Structure Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers South Pasadena Bonding And Molecular Structure Worksheet Answers
9. Balancing eBooks and Physical Books South Pasadena Bonding And Molecular Structure Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection South Pasadena Bonding And Molecular Structure Worksheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine South Pasadena Bonding And Molecular Structure Worksheet Answers
 - Setting Reading Goals South Pasadena Bonding And Molecular Structure Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of South Pasadena Bonding And Molecular Structure Worksheet Answers
 - Fact-Checking eBook Content of South Pasadena Bonding And Molecular Structure Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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