



ACIDS & BASES 1

- 1) a) What is a base? (1)
b) What is an alkali? (1)
c) What is the reaction between an acid and a base called? (1)
d) Give two everyday examples of reactions between acids and bases.
1
2 (2)

- 2) a) Match the descriptions below to the pH's in this list: 2 5 7 9 13 (2)

description	neutral	very alkaline	slightly acidic	slightly alkaline	very acidic
pH					

- b) i) Litmus indicator could not distinguish between solutions of pH 2 and 5. Explain why. (1)
ii) Name an indicator that could. (1)

- 3) Complete the following word equations: .
a) zinc + nitric acid $\rightarrow$ (2)
b) hydrochloric acid + copper carbonate $\rightarrow$ (2)
c) sulfuric acid + sodium oxide $\rightarrow$ (2)
d) potassium hydroxide + sulfuric acid $\rightarrow$ (2)
e) hydrochloric acid + lithium carbonate $\rightarrow$ (2)
f) magnesium oxide + nitric acid $\rightarrow$ (2)
g) sulfuric acid + strontium $\rightarrow$ (2)
h) potassium carbonate + sulfuric acid $\rightarrow$ (2)
i) rubidium hydroxide + hydrochloric acid $\rightarrow$ (2)

- 4) a) Why should you not add an alkali to an acid spill? (2)
b) Give two safer ways of treating acids spills.
1
2 (2)

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