

8.2

THE NATURE OF COVALENT BONDING

KEY

Section Review

Objectives

- State a rule that usually tells how many electrons are shared to form a covalent bond
- Describe how electron dot formulas are used
- Predict when two atoms are likely to be joined by a double or a triple covalent bond
- Distinguish between a single covalent bond and other covalent bonds
- Describe how the strength of a covalent bond is related to its bond dissociation energy
- Describe how resonance structures explain bonding
- Identify some exceptions to the octet rule

Vocabulary

- single covalent bond
- structural formulas
- unshared pairs
- double covalent bonds
- triple covalent bonds
- coordinate covalent bond
- polyatomic ion
- bond dissociation energy
- resonance structures

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

When atoms share electrons to gain the 1 configuration of a noble gas, the bonds formed are 2. A 3 pair of valence electrons constitutes a 4 covalent bond. Pairs of valence electrons that are not shared between atoms are called 5. Sometimes two or three pairs of electrons may be shared to give 6 covalent bonds. In some cases, only one of the atoms in a bond provides the pair of bonding electrons; this is a 7. 8 is required to break covalent bonds between atoms. The total energy required to break the bond between two covalently bonded atoms is known as the 9.

When it is possible to write two or more valid electron dot formulas for a molecule or ion, each formula is referred to as a 10.

1. stable
2. covalent
3. shared
4. single
5. unshared pairs
6. double/triple
7. coordinate covalent bond
8. energy
9. bond dissociation energy
10. resonance structure

The Nature Of Covalent Bonding 8 2 Answer Key

Fox



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Chemistry Martin Stuart Silberberg, 1996 This general chemistry text centres on the theme that observable change in chemical systems is the result of molecular change The aims of this edition are to enable students to perceive matter and change at the molecular level and to help build student confidence in their ability to solve chemical problems as they discover the relevance of chemistry to their lives

Prentice Hall Chemistry, 2000 Basic Concepts of Chemistry Leo J. Malone, Theodore O. Dolter, 2012-04-13 Basic Concepts of Chemistry was originally written over thirty years ago to address the needs of general chemistry students

with little or no background in chemistry Over time the text has evolved beyond purposes solely aimed at a preparatory chemistry course For some preparatory chemistry students a main sequence in general chemistry may follow but for others a semester of organic and biochemistry may follow Other students enroll to simply satisfy a basic science or chemistry requirement The text was written with a level and functionality designed to accommodate the needs of each of these varied groups of students Basic Concepts of Chemistry was designed with a flexibility that allows instructors to emphasize or omit certain clearly delineated sections The mission of the text has evolved in response to the increased diversity of students and the emphasis on outcomes assessment

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Table of Contents The Nature Of Covalent Bonding 8 2 Answer Key

1. Understanding the eBook The Nature Of Covalent Bonding 8 2 Answer Key
 - The Rise of Digital Reading The Nature Of Covalent Bonding 8 2 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying The Nature Of Covalent Bonding 8 2 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an The Nature Of Covalent Bonding 8 2 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from The Nature Of Covalent Bonding 8 2 Answer Key

- Personalized Recommendations
 - The Nature Of Covalent Bonding 8 2 Answer Key User Reviews and Ratings
 - The Nature Of Covalent Bonding 8 2 Answer Key and Bestseller Lists
5. Accessing The Nature Of Covalent Bonding 8 2 Answer Key Free and Paid eBooks
 - The Nature Of Covalent Bonding 8 2 Answer Key Public Domain eBooks
 - The Nature Of Covalent Bonding 8 2 Answer Key eBook Subscription Services
 - The Nature Of Covalent Bonding 8 2 Answer Key Budget-Friendly Options
 6. Navigating The Nature Of Covalent Bonding 8 2 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - The Nature Of Covalent Bonding 8 2 Answer Key Compatibility with Devices
 - The Nature Of Covalent Bonding 8 2 Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of The Nature Of Covalent Bonding 8 2 Answer Key
 - Highlighting and Note-Taking The Nature Of Covalent Bonding 8 2 Answer Key
 - Interactive Elements The Nature Of Covalent Bonding 8 2 Answer Key
 8. Staying Engaged with The Nature Of Covalent Bonding 8 2 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers The Nature Of Covalent Bonding 8 2 Answer Key
 9. Balancing eBooks and Physical Books The Nature Of Covalent Bonding 8 2 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection The Nature Of Covalent Bonding 8 2 Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine The Nature Of Covalent Bonding 8 2 Answer Key
 - Setting Reading Goals The Nature Of Covalent Bonding 8 2 Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of The Nature Of Covalent Bonding 8 2 Answer Key

- Fact-Checking eBook Content of The Nature Of Covalent Bonding 8 2 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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