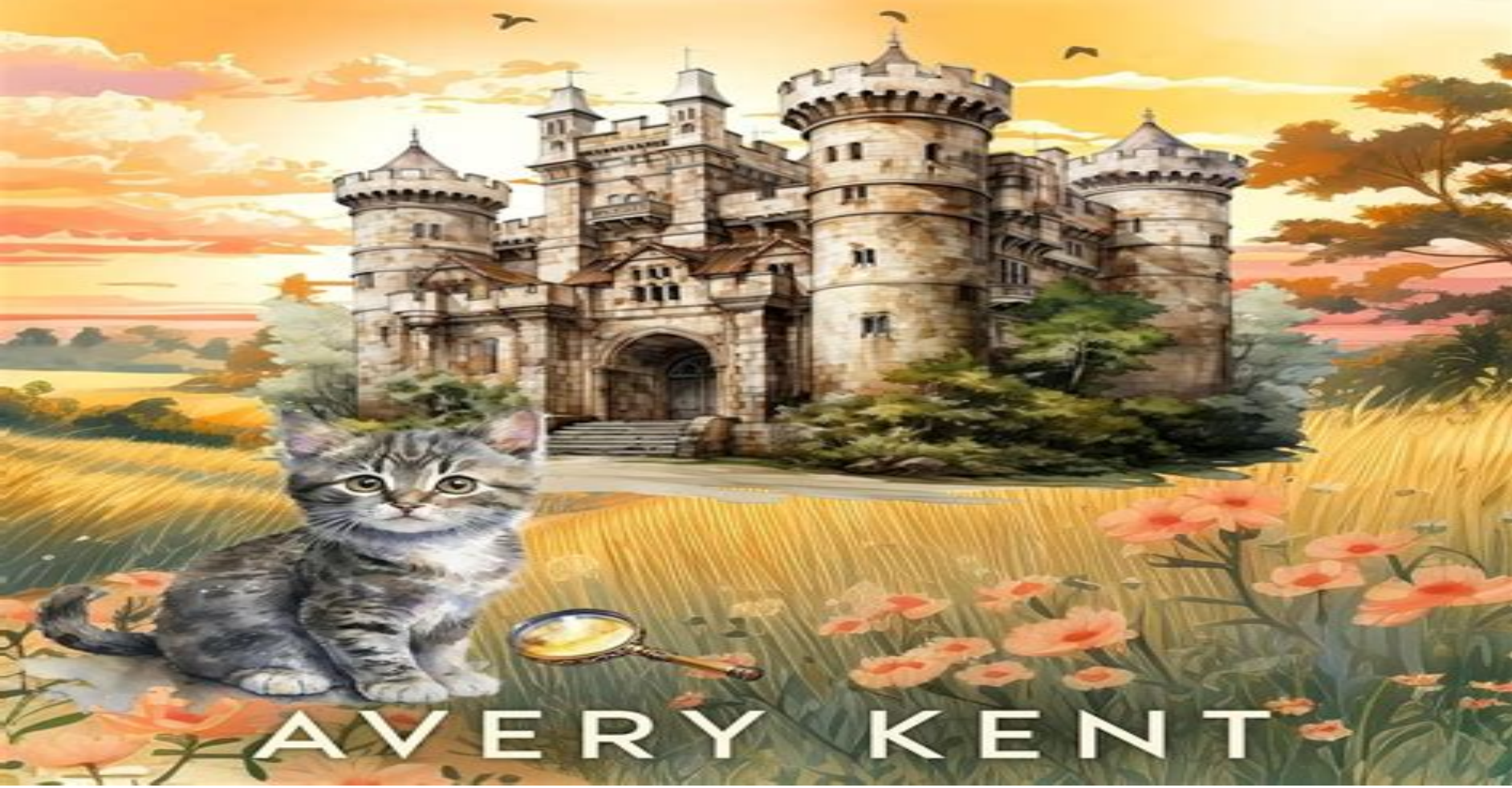


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 temperature change plugging in the values given in the problem you get $q_{\text{water}} = 4.18 \text{ J/g} \times 110 \text{ g} \times 26.6 \text{ C} - 25.0 \text{ C}$ $q_{\text{water}} = 550 \text{ J}$ $\Delta h_{\text{q water}} = 550 \text{ J}$

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 $14.73 \text{ kJ} / 14.73 \text{ mol} = 1 \text{ kJ/mol}$

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web 6 a 218 g sample of steam at 121 °C is cooled to ice at 14 °C find the change in heat content of the system 7 in going from ice at 34 °C to steam at 138 °C a sample of water absorbs 1.41×10^5 J find the mass of the sample answers 1 2 68 x 10⁴ J 2 6 23 x 10⁴ J 3 1 71 x 10⁴ J 4 1 11 x 10⁵ J 5 2 11 x 10⁵ J 6

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