

12 Stoichiometry



THE MOLE AND QUANTIFYING MATTER, REACTIONS

12.1 The Arithmetic of Equations

Learning Objectives

By the end of this section, students should be able to apply the arithmetic of equations.

Learning Objectives

Having completed this section, students should be able to:

1. Balance the equation of an unbalanced reaction, allowing mass conservation to be observed.

Chemical Equations: The equations start with formulae (subscripts and balanced chemical equations).

1. Read balanced chemical equations: the total number of atoms of each element before reaction must equal the total number of atoms of each element after reaction.
2. Read balanced chemical equations: the total mass of the reactants must equal the total mass of the products.

After reading section 12.1, answer the following questions.

Learning Objectives

1. Balance the equation of an unbalanced reaction, allowing mass conservation to be observed.

2. Identify correctly the number _____ of molecules present in a given reaction.

3. In balancing the total mass of atoms (H, O, C, N, S, Fe, etc.) must be the same on both sides of a balanced equation. But there is something wrong with this.

4. In the following unbalanced reaction, the number of atoms of oxygen is different. Why? _____

5. Calculate the mass of the products in the reaction: _____

Chemical Equations

6. If you take 100g of reactants, how much product will you get? _____

The Mole And Quantifying Matter Reactions Answers

Camilla Rothe



The Mole And Quantifying Matter Reactions Answers:

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