

Sample Problem 1:

Calculate the *Molarity* of a solution made by dissolving 2.355 g of copper (II) acetate (FW 181.5) in 50.0 mL of water.

$$0.259504 \text{ M} = 0.260 \text{ M}$$

KEY

$$\frac{2.355 \text{ g Cu}(\text{C}_2\text{H}_3\text{O}_2)_2}{50.0 \text{ mL soln}} \times \frac{1 \text{ mol Cu}(\text{C}_2\text{H}_3\text{O}_2)_2}{181.5 \text{ g Cu}(\text{C}_2\text{H}_3\text{O}_2)_2} \times \frac{1000 \text{ mL soln}}{1 \text{ L soln}} = 0.259504 \text{ M} = 0.260 \text{ M}$$

38F

Sample Problem 2:

Calculate the *concentration in ppm, pp thousand* of a solution made by dissolving 2.355 g of copper (II) acetate (FW 181.5) and diluting up with water in a 50.0 mL volumetric flask. The density of solution at the given temperature is 1.0013 g/mL.

$$4.70388 \times 10^4 \text{ ppm} ; 4.70 \times 10^4 \text{ ppth}$$

$$\frac{2.355 \text{ g Cu}(\text{C}_2\text{H}_3\text{O}_2)_2}{50.0 \text{ mL soln}} \times \frac{1 \text{ mL soln}}{1.0013 \text{ g soln}} = \frac{x}{1.0 \times 10^6}$$

$\frac{\text{g solute}}{\text{g soln}} = \frac{x}{1 \times 10^6}$

$$\frac{0.0470388 \text{ g Cu}(\text{C}_2\text{H}_3\text{O}_2)_2}{1 \text{ g soln}} = \frac{x}{1.0 \times 10^6} ; \quad x = 47,000 \text{ ppm}$$

OR
47.0388 ppthousand

Sample Problem 3:

How many grams of solute would be used to prepare 1.50 L of a 0.250 M NaOH solution?

$$1.50 \text{ L NaOH soln} \times \frac{0.250 \text{ mol NaOH}}{1 \text{ L NaOH soln}} \times \frac{40.0 \text{ g NaOH}}{1 \text{ mol NaOH}}$$

$$= 15.0 \text{ g NaOH}$$

dissolved in approximately

0.75 L, then

brought up to total Volume of 1.52

$$\begin{array}{r} \text{Na} \quad 23 \\ \text{O} \quad 16 \\ \text{H} \quad 1 \\ \hline 40.0 \text{ g} \\ \text{mole} \end{array}$$

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