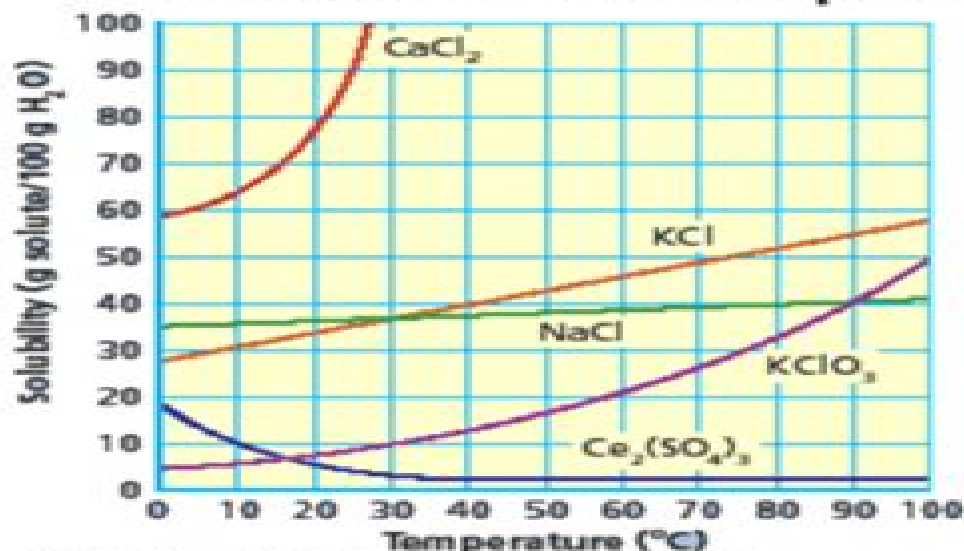


Directions: Use the graph below to answer the following questions. If the question requires a calculation, SHOW ALL WORK TO RECEIVE CREDIT FOR THE ANSWER. No work, no credit, no kidding! Remember units.

### Solubilities as a Function of Temperature



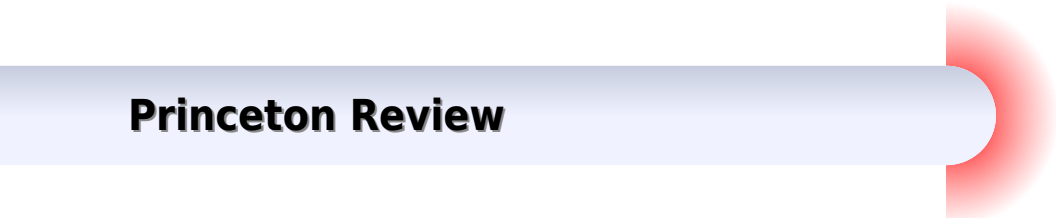
#### Types of Solutions

- Points that collectively make up the solubility curve (points ON the curve) represent \_\_\_\_\_ solutions.
- Points that are BELOW the curve represent \_\_\_\_\_ solutions.
- Points ABOVE the solubility curve represent \_\_\_\_\_ solutions, and the difference between the point above the curve and on the curve represents the amount of solute which will precipitate out.

1. a. What is the solubility of calcium chloride at 5°C? \_\_\_\_\_g/100g of H<sub>2</sub>O  
 b. What is the solubility of calcium chloride at 25°C? \_\_\_\_\_g/100g of H<sub>2</sub>O  
 c. What is happening to the solubility and temperature?  
 Solubility is \_\_\_\_\_Increasing or Decreasing\_\_\_\_\_ as temperature is \_\_\_\_\_Increasing or Decreasing\_\_\_\_\_.  
 What is this type of relationship called? \_\_\_\_\_direct or indirect\_\_\_\_\_.  
 This occurs for which type of solutes? \_\_\_\_\_gasses or solids\_\_\_\_\_.
2. a. What is the solubility of cerium sulfate at 10°C? \_\_\_\_\_g/100g of H<sub>2</sub>O  
 b. What is the solubility of cerium sulfate at 50°C? \_\_\_\_\_g/100g of H<sub>2</sub>O  
 c. What is happening to the solubility and temperature?  
 Solubility is \_\_\_\_\_increasing or decreasing\_\_\_\_\_ as temperature is \_\_\_\_\_increasing or decreasing\_\_\_\_\_.  
 What is this type of relationship called? \_\_\_\_\_direct or indirect\_\_\_\_\_.
3. a. At 90°C, 10 g of potassium chlorate is dissolved in 100. g of water.  
 Is this solution saturated, unsaturated, or supersaturated?  
 b. How do you know?

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**Princeton Review**



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