# Table F Solubility Guidelines for Aqueous solutions

Ions That Form Soluble Compounds	Exceptions	Ions Th Insoluble C	
Group 1 ions (Li+, Na+, etc.)		carbonate (	
ammonium (NH <sub>4</sub> +)		chromate (	
nitrate (NO <sub>3</sub> <sup>-</sup> )		11	
acetate (C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup> or CH <sub>3</sub> COO <sup>-</sup> )		phosphate (	
hydrogen carbonate (HCO <sub>3</sub> <sup>-</sup> )		sulfide (S <sup>2</sup> -	
chlorate (ClO <sub>3</sub> ~)		hydroxide (	
halides (Cl <sup>+</sup> , Br <sup>+</sup> , I <sup>+</sup> )	when combined with Ag*, Pb <sup>2</sup> *, or Hg <sub>2</sub> <sup>2</sup> *		
sulfates (SO <sub>4</sub> <sup>2-</sup> )	when combined with Ag*, Ca <sup>2+</sup> , Sr <sup>2+</sup> , Ba <sup>2+</sup> , or Pb <sup>2+</sup>	*compounds l	

Ions That Form Insoluble Compounds*	Exceptions	
carbonate (CO <sub>3</sub> <sup>2-</sup> )	when combined with Group I fons or ammonium $(\mathrm{NH_4}^+)$	
chromate (CrO <sub>4</sub> <sup>2-</sup> )	when combined with Group I ions, $Ca^{2+}$ , $Mg^{2+}$ , or ammonium $(NH_4^+)$	
phosphate (PO <sub>4</sub> <sup>3-</sup> )	when combined with Group I tons or ammonium (NH <sub>4</sub> *)	
sulfide (S <sup>2-</sup> )	when combined with Group 1 ions or ammonium (NH <sub>4</sub> *)	
hydroxide (OH*)	when combined with Group 1 ions, Ca <sup>2+</sup> , Ba <sup>2+</sup> , Sr <sup>2+</sup> , or ammonium (NH <sub>4</sub> *)	

<sup>\*</sup>compounds having very low solubility in H<sub>2</sub>O

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**Richard Bailey** 

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