Student Solution Manual

Student Solution Manual: Skoog et al., *Fundamentals of Analytical Chemistry*, 10e,
© 2022, 978-0-357-45055-0, Chapter 2: Calculations Used in Analytical Chemistry

Some of the answers below may differ in format but have the same value as your result. Please check with your instructor if a specific format is desired.

Chapter 2

2-1. Define

Answers:

(a) molar mass.

The molar mass is the mass in grams of one mole of a chemical species.

(c) millimolar mass.

The *millimolar mass* is the mass in grams of one millimole of a chemical species.

2-3. Give two examples of units derived from the fundamental base SI units.

Solution:

$$\text{The liter: } 1\text{ L} = \frac{1000\text{ mL}}{1\text{ L}} \times \frac{1\text{ cm}^3}{\text{mL}} \times \left(\frac{\text{m}}{100\text{ cm}}\right)^3 = 10^{-3}\text{ m}^3$$

$$\text{Molar concentration: } 1\text{ M} = \frac{1\text{ mol}}{\text{L}} \times \frac{\text{L}}{10^{-3}\text{ m}^3} = \frac{1\text{ mol}}{10^{-3}\text{ m}^3}$$

2-4. Simplify the following quantities using a unit with an appropriate prefix:

Solutions:

(a) 5.8×10^8 Hz.

$$5.8 \times 10^8\text{ Hz} \times \frac{\text{MHz}}{10^6\text{ Hz}} = 580\text{ MHz}$$

(c) 9.31×10^7 μmol.

$$9.31 \times 10^7\ \mu\text{mol} \times \frac{\text{mol}}{10^6\ \mu\text{mol}} = 93.1\text{ mol}$$

(e) 3.96×10^5 nm.

$$3.96 \times 10^5\text{ nm} \times \frac{\text{mm}}{10^6\text{ nm}} = 3.96\text{ mm}$$

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Sung-Hyuk Cha

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