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* Universal gas constant $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$ * Avogadro constant $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$

PART B - ESSAY

Answer two questions only. (Each question carries 150 marks.)

(a) Consider the reaction (1) given below at a temperature of 800 °C.

 $H_3(g) + I_2(g) \implies 2HI(g)$ -----(1)

Initially, 0.45 mol of HI(g) was introduced into a 1.0 dm³ evacuated rigid closed container at 800 °C, and allowed to reach the equilibrium given above. It was found that 0.05 mol of $H_2(g)$ was present at equilibrium.

- (i) Calculate the equilibrium constant K_{C1}, for the above equilibrium at a temperature of 800 °C.
- (ii) In a separate similar evacuated container at a temperature of 800 °C, reaction (2) takes place with an equilibrium constant K_{C2} = 1.2 × 10⁸ mol⁻¹ dm³.

 $2H_2(g) + S_2(g) \Rightarrow 2H_2S(g) - (2)$

When the two containers are connected together, the following reaction (3) takes place at a temperature of 800 °C.

 $2I_2(g) + 2H_2S(g) \implies S_2(g) + 4HI(g)$ -----(3)

Calculate the equilibrium constant K_{C_3} for the reaction (3) at a temperature of 800 °C.

- (iii) A 1.0 dm³ rigid closed container at a temperature of 800 °C contains an equilibrium mixture of (3) in (ii) above with 5.00 × 10⁻⁵ mol HI(g), 1.25 × 10⁻⁶ mol S₂(g) and 2.50 × 10⁻⁵ mol H₂S(g). Calculate the number of moles of I₂(g) in the above mixture.
- (iv) An extra 2.50×10^{-3} mol $I_2(g)$ was added to the equilibrium mixture in (iii) above at a temperature of 800 °C.
 - Calculate the reaction quotient (Q_C) at the moment the extra I₂(g) is added.
 - II. Explain the change in the equilibrium that takes place upon the addition of extra I2(g).
 - III. Sketch the variation in the concentrations of each constituent in the mixture, with time, upon the addition of extra I₂(g).
 (60 marks)
- (b) (i) Using the data given below, calculate ΔH°, ΔS° and ΔG° for the reaction (4) at 27 °C.

 $H_2(g) + I_2(g) \longrightarrow 2HI(g)$ -----(4)

At 27 °C: $H_2(g) + I_2(s) \rightarrow 2 HI(g)$; $\Delta H^0 = 53 \text{ kJ mol}^{-1}$, $\Delta S^0 = 410 \text{ J K}^{-1} \text{ mol}^{-1}$ $I_2(s) \rightarrow I_2(g)$; $\Delta H^0 = 63 \text{ kJ mol}^{-1}$, $\Delta S^0 = 260 \text{ J K}^{-1} \text{ mol}^{-1}$

(ii) Using the data given below, calculate ΔH^* , ΔS^* and ΔG^* for the reaction (5) at 27 °C. $2 \, \text{H}_2 \text{S}(g) \longrightarrow 2 \, \text{H}_2(g) + S_2(g) ------(5)$

[see page ten

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T. A. Burton

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