Name: KEY

Valence Electrons Practice

The valence electrons are the electrons in the outermost energy level. They are always the outermost electrons.

Atoms that LOSE electrons are called cations and have an overall positive charge.

Atoms that GAIN electrons are called anions and have an overall negative charge.

Determine the element's number of valence electrons (# of electrons in the "outermost" shell).

Example: Carbon has 4 valence electrons, carbon 4

- 1. fluorine 7
 2. lithium 1
 3. phosphorus 5
 4. francium 2
 5. calcium 2
 6. carbon 4
 7. nitrogen 5
- 8. iodine / 9. silicon + 10. oxygen 6 11. argon 5 12. barium 2 13. potassium 1 14. aluminum 3

15. helium 2.

16. hydrogen 1.

17. magnesium 2.

18. xenon 3.

19. sulfur (...

20. boron 3.

21. bromine 7.

22. What are valence electrons of an atom used for?

Octermining how elements bond together

- 23. Which groups on the periodic table want to lose electrons? What kind(s) of elements are these groups? Group 1,2 elements, Alk; metals and Alk; rarth
- 24. Which groups on the periodic table want to gain electrons? What kind(s) of elements are these groups? Groups 5. (, 7, non-molals
- 25. Why do the elements from question 23 want to lose electrons? Why do the elements from question 24 want to gain electrons? (HINT: The answer is the same for both questions.).

 By laring 1 or 2 electrons then groups form prove stable ion having the electron configuration of noble gases
- 26. How many valence electrons do the Alkaline Earth metals have? Are these elements very reactive? Explain your answer. 2 , ye s
- 27. What is the name of the atoms that lose electrons? What is their overall charge?

 Cofinn(4) charge
- 28. What is the name of the atoms that gain electrons? What is their overall charge?

 Anions (-) charge

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