

- (c) 350 cm, convergent
(d) 150 cm, divergent
16. Three resistances A, B and C each of 5Ω are connected to a battery of 10 V as shown in the figure. Calculate the current through C:
- 
- (a) 2 amp (b) 1 amp
(c) 3 amp (d) 0.6 amp
17. The device used to generate electrical energy is:
- (a) Electric motor (b) Generator
(c) Galvanometer (d) Voltmeter
18. The number of molecules in 11 g of CO_2 are:
- (a) 0.25×10^{23} (b) 0.50×10^{23}
(c) 1.00×10^{23} (d) 1.51×10^{23}
19. The value of charge / mass ratio of electron was determined by:
- (a) W.K. Roentgen (b) J.J. Thomson
(c) Marie Curie (d) Niels Bohr
20. Which of the following reactions is an example of combination reaction?
- (a) $\text{Zn(s)} + \text{CuSO}_4(\text{aq}) \rightarrow \text{ZnSO}_4(\text{aq}) + \text{Cu(s)}$
(b) $\text{BaCl}_2(\text{aq}) + \text{Na}_2\text{SO}_4(\text{aq}) \rightarrow \text{BaSO}_4(\text{s}) + 2\text{NaCl}(\text{aq})$
(c) $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO(s)} + \text{CO}_2(\text{g})$
(d) $\text{C(s)} + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$
21. The element 'X' and 'Y' have atomic numbers 12 and 17 respectively. Element 'X' reacts with element 'Y' to form compound with molecular formula:
- (a) XY (b) XY_2
(c) XY_3 (d) X_3Y_3
22. The atomic radius (pm) of Li, Na, K and Rb varies in the order:
- (a) $\text{Na} < \text{K} < \text{Rb} < \text{Li}$ (b) $\text{K} < \text{Na} < \text{Li} < \text{Rb}$
(c) $\text{Li} < \text{Na} < \text{K} < \text{Rb}$ (d) $\text{Rb} < \text{K} < \text{Na} < \text{Li}$
23. The electron affinity value (kJ mol^{-1}) of fluorine (F) is less than:
- (a) Hydrogen (H) (b) Lithium (Li)
(c) Oxygen (O) (d) Chlorine (Cl)
24. Which of the following is an example of strong electrolyte?
- (a) H_2CO_3 (b) NH_4OH
(c) NaCl (d) $\text{COOH} - \text{COOH}$
25. If pH of a solution changes from 5 to 4, the change in hydrogen ion concentration shall be:
- (a) two times (b) five times
(c) ten times (d) twenty times
26. Bleaching powder is manufactured by the reaction of:
- (a) CaCl_2 and CaCO_3 (b) Cl_2 and Ca(OH)_2
27. (c) Cl_2 and CaSO_4 (d) Cl_2 and $\text{Ca(HCO}_3)_2$
The hard glass is obtained by fusing:
- (a) soda ash, sand and limestone
(b) a mixture of sand, lime, borax and alkali carbonates
(c) potassium carbonate and limestone
(d) potassium carbonate, lead oxide and sand
28. Which of the following metals is most reactive?
- (a) Aluminium (b) lead
(c) Mercury (d) Silver
29. Which of the following alloys contains chromium?
- (a) Steel (b) Stainless steel
(c) Magnadium (d) Brass
30. Which of the following is monomer of natural rubber?
- (a) Chloroethene (b) Chloroprene
(c) Isoprene (d) Buta-1, 3-diene
31. Alkaline KMnO_4 oxidises propanone to:
- (a) propanoic acid (b) ethanoic acid
(c) methanoic acid (d) oxalic acid
32. Which of the following metals can displace Zn from ZnSO_4 solution?
- (a) Calcium (b) Copper
(c) Iron (d) Mercury
33. Heating of sodium ethanoate with soda lime yields:
- (a) ethane (b) methane
(c) ethanol (d) methanol
34. The compound formed by the reaction of ethyne with bromine is:
- (a) $\text{Br}-\text{CH}=\text{CH}-\text{Br}$ (b) $\text{Br}-\text{CH}_2-\text{CHBr}_2$
(c) $\text{CH}_2=\text{CH}-\text{Br}$ (d) $\text{Br}_2\text{CH}-\text{CHBr}_2$
35. Cell were first discovered by Robert Hooke in the year:
- (a) 1665 (b) 1674 (c) 1831 (d) 1839
36. The plant cells have a rigid cell wall that lies:
- (a) Outside the plasma membrane
(b) Inside the plasma membrane
(c) In between the plasma membranes
(d) None of the above
37. During mitosis nucleolus and nuclear membrane are lost in stage:
- (a) Prophase (b) Metaphase
(c) Anaphase (d) Telophase
38. An example of simple and permanent tissue is:
- (a) Xylem (b) Phloem
(c) Sclerenchyma (d) All the above
39. The shape of squamous epithelial tissue is:
- (a) Cubical (b) Flattened
(c) Pillar like (d) None of the above
40. Fresh water sponge belongs to phylum:
- (a) Arthropoda (b) Annelida

Viswavarati University Entrance Question Paper

**Pradipta Maji, Tingwen Huang, Nikhil R.
Pal, Santanu Chaudhury, Rajat K. De**

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Table of Contents Viswavarati University Entrance Question Paper

1. Understanding the eBook Viswavarati University Entrance Question Paper
 - The Rise of Digital Reading Viswavarati University Entrance Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Viswavarati University Entrance Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Viswavarati University Entrance Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Viswavarati University Entrance Question Paper
 - Personalized Recommendations
 - Viswavarati University Entrance Question Paper User Reviews and Ratings
 - Viswavarati University Entrance Question Paper and Bestseller Lists
5. Accessing Viswavarati University Entrance Question Paper Free and Paid eBooks
 - Viswavarati University Entrance Question Paper Public Domain eBooks
 - Viswavarati University Entrance Question Paper eBook Subscription Services
 - Viswavarati University Entrance Question Paper Budget-Friendly Options
6. Navigating Viswavarati University Entrance Question Paper eBook Formats

- ePub, PDF, MOBI, and More
- Viswavarati University Entrance Question Paper Compatibility with Devices
- Viswavarati University Entrance Question Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Viswavarati University Entrance Question Paper
 - Highlighting and Note-Taking Viswavarati University Entrance Question Paper
 - Interactive Elements Viswavarati University Entrance Question Paper
- 8. Staying Engaged with Viswavarati University Entrance Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Viswavarati University Entrance Question Paper
- 9. Balancing eBooks and Physical Books Viswavarati University Entrance Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Viswavarati University Entrance Question Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Viswavarati University Entrance Question Paper
 - Setting Reading Goals Viswavarati University Entrance Question Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Viswavarati University Entrance Question Paper
 - Fact-Checking eBook Content of Viswavarati University Entrance Question Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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