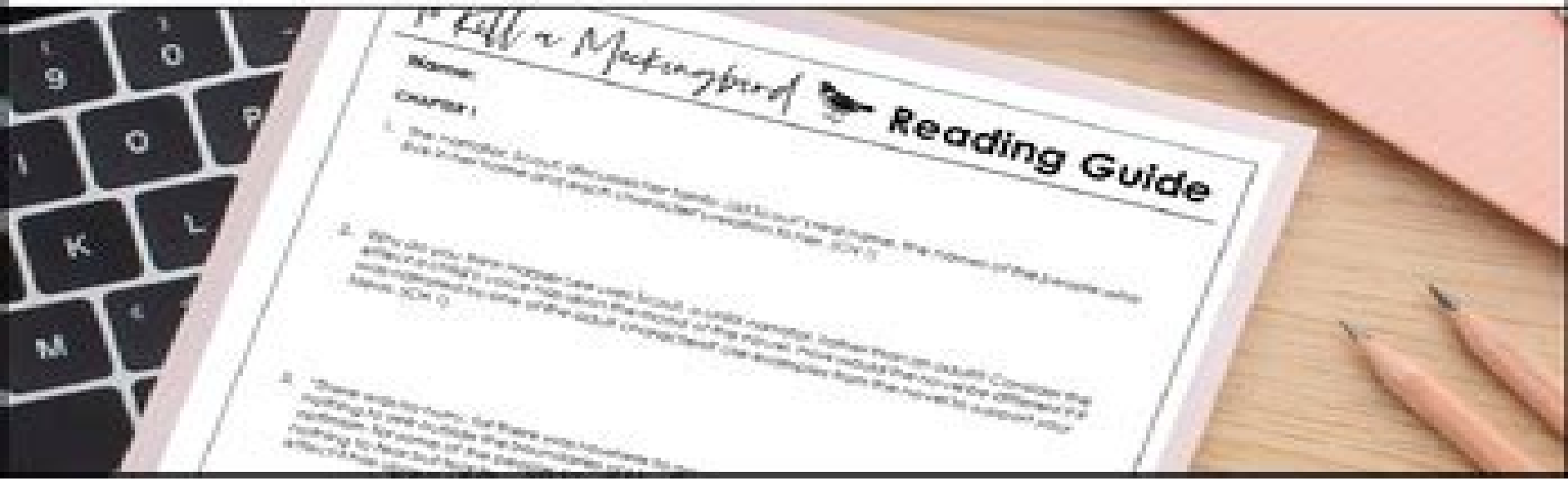


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diagram may depend on the version and year of production of x3 this one is for 2004 2006 gb version

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 $2 \text{ CO}_2(\text{g}) \rightarrow 2 \text{ C}(\text{s}) + 2 \text{ O}_2(\text{g}) \quad \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) \rightarrow 2 \text{ C}(\text{s}) + 3 \text{ H}_2(\text{g}) \quad \Delta h = 283 \text{ kJ}$

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combustion of magnesium oxide mg s 1 2o2 g mgo s we used

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web thus for reaction 9 2 in which 1 mole of h_3o and 1 mole of oh combine to form 2 moles of h_2o $\Delta h = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta h_{products} - \Delta h_{reactants}$ is positive if heat is absorbed that is if $\Delta h_{products}$

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 $\text{1 NaOH (s)} + \text{Na}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$
 $\text{2 NaOH (s)} + \text{H}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$
 $\text{3 NaOH (s)} + \text{H}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$