

6.9) It is often useful to write the equations for the reactions described:



The reaction for which the heat of reaction is requested is the formation of methane from its constituent elements.



The ΔH for this reaction can be calculated by summing the ΔH 's for a series of reactions which together add up to the same overall reaction (Hess's Law of Heat Summation).



These equations add up to the desired equation, so the ΔH values are added:

$$\Delta H^\circ(\text{CH}_4) = +890. - 394 - 572 \text{ kJ/mol} = -76 \text{ kJ/mol}$$

6.10) $\Delta H_{\text{vap}} = -924 \text{ kJ}$

6.11) The heat of reaction is the difference between the heat of formation of the products and the heat of formation of the reactants in the balanced equation:

$$\Delta H^\circ = 12 \times (-384) + 6 \times (-242) - 2 \times (+49) = -8278 \text{ kJ.}$$

There are two moles of benzene in the equation, so the heat released for one mole is 1/2 this amount, 3139 kJ/mol.

6.12) $\Delta H_{\text{vap}} = -95.6 \text{ kJ}$

6.13) a. Heat lost = $50. \text{ g} \times 0.129 \text{ J/g}^\circ\text{C} \times 68^\circ\text{C} = 260 \text{ J}$

b. $q(\text{ice melting}) + q(\text{water warming}) = -q(\text{water cooling})$

$$5.0 \text{ g} \times 334 \text{ J/g} + 5.0 \text{ g} \times 4.184 \text{ J/g}^\circ\text{C} (T_f - 0^\circ\text{C}) = -100 \text{ g} \times 4.184 \text{ J/g}^\circ\text{C} (T_f - 70^\circ\text{C})$$

$$1670 \text{ J} + 20.92 T_f = -418.4 T_f + 29,288 \text{ J}$$

$$439.32 T_f = 27,618$$

$$T_f = 62.9 = 63^\circ\text{C}$$

6.14) a. $q(\text{water}) = -q(\text{coins})$

$$c \times m \times \Delta T = -c \times m \times \Delta T$$

$$4.184 \text{ J/g}^\circ\text{C} \times 200. \text{ g} \times 5.9^\circ\text{C} = -c \times 113.4 \text{ g} \times (-74.1^\circ\text{C})$$

$$\text{Solving for } c \text{ gives: } 0.587 \text{ J/g}^\circ\text{C}$$

b. No, the specific heat capacity of pure gold is $0.129 \text{ J/g}^\circ\text{C}$.

6.15) $\Delta H = 4.18 \text{ J/g}^\circ\text{C} \times 22.4 \text{ K} \times 70. \text{ g} = 6560 \text{ J}$

$$\text{For one mole of ammonium nitrate, } \Delta H = 6560 \times 60.05/20 = 26 \text{ kJ/mol.}$$

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web l étudiant peut choisir parmi les 5 grands types de logement étudiant les petites annonces de particuliers à particuliers sans frais d agence les résidences étudiantes privées les chambres chez l habitant la colocation et les cités u exclusivement dédié à la location de petites surfaces les particuliers bailleurs peuvent

logement étudiant vie étudiante l étudiant - Aug 16 2023

web logement étudiant les années étudiantes sont souvent marquées par le premier appartement loin des parents choix de la location budget à y consacrer aides

accommodation Étudiant gouv - May 01 2022

web dans une résidence universitaire dans un logement privé loué par une agence immobilière ou un particulier il y a différentes façons de trouver un logement en france accommodation Étudiant gouv

location appartement meublé étudiant louer studio appartement - Dec 28 2021

web location d appartement meublé pour les étudiants louer un appartement un studio ou une chambre meublée dans une résidence hôtelière accueillant les étudiants sur appart city leader de l appart hôtel en france