

Names: _____

Thermodynamics IN CLASS Skill Review

Skill - Can define and identify direction of heat flow for exothermic and endothermic situations. Before completing the calculations below, answer the following:

- Provide the labels exothermic and endothermic, to the following pictures:

a. What are some everyday examples of

endothermic:

exothermic and endothermic reactions?

Answers may vary - exothermic (hot pad),

endothermic (baking a cake).

exothermic:



- Label the following reactions as being exothermic or endothermic:

a. As a tree grows, photosynthesis **absorbs** energy from the environment to break apart CO_2 and H_2O .

endothermic

b. Sweating cools a person down as energy is **released** from the skin to change water into a gas.

exothermic

Skill - Use the equation $\Delta E = q + w$ correctly when given these terms in a real-world situation.

For the following, calculate the ΔE and label each scenario as describing an exothermic or endothermic reaction:

- A system releases 23 J of heat while 12 J of work is done on it. What is the ΔE ?

$$\Delta E = q + w \rightarrow -23\text{J} + 12\text{J} = \mathbf{-11\text{J} \text{ exothermic}}$$

- If a system absorbs 67 J of heat and performs 32 J of work, what is the ΔE ?

$$\Delta E = q + w \rightarrow 67\text{J} + -32\text{J} = \mathbf{35\text{J} \text{ endothermic}}$$

Skill - Use $q = m \cdot c \cdot \Delta T$ to calculate the missing quantity, given the others.

Be sure to show all work required for the following calculations:

- Graphite has specific heat of $0.709 \text{ J/g}^\circ\text{C}$. If a 25.0g piece of graphite is cooled from 35°C to 18°C , how much energy was lost by the graphite? Is the reaction endothermic or exothermic?

$$q = m \cdot c \cdot \Delta T \rightarrow 0.709\text{J/g}^\circ\text{C} \times 25.0\text{g} \times -17^\circ\text{C} = \mathbf{-300\text{J} \text{ exothermic}}$$

- An unknown substance with a mass of 100.0g, **absorbs** $1.0 \times 10^3 \text{ J}$ while undergoing a temperature increase of 25°C . What is the specific heat of the substance? Is the reaction endothermic or exothermic?

$$q = m \cdot c \cdot \Delta T \rightarrow c = q / (m \cdot \Delta T) \rightarrow 1.0 \times 10^3 \text{ J} / (100.0\text{g} \times 25^\circ\text{C}) = \mathbf{0.4\text{J/g}^\circ\text{C}} \text{, } q \text{ is endothermic}$$

Smurray Electricity And Thermodynamics Review Answer Key

Christian G. Meyer



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