

Student Exploration: Titration

	Get the Gizmo ready:	
Activity A: Acids and bases	Click Reset. Select 1.0 M HNO, for the Burette and Mystery NaOH for the Flask. Select Phenolphthalein for the Indicator. You will need a scientific calculator for this activity.	ar post
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Question: How do acids and bases interact in solution?

- <u>Describe</u>: The equation for the reaction of nitric acid (HNO₃) and sodium hydroxide (NaOH) is shown on the bottom right of the Gizmo.
 - A. What are the reactants in this reaction? HNO₃, NaOH
 - B. What are the products of this reaction? NaNO₃, H₂O
- Measure: A titration can be used to determine the concentration of an acid or base by
 measuring the amount of a solution with a known concentration, called the titrant, which
 reacts completely with a solution of unknown concentration, called the analyte. The point at
 which this occurs is called the equivalence point.

Carefully add HNO₃ into the flask until the phenolphthalein begins to lose its color. Stop adding HNO₃ when the color change is permanent.

- A. How much (HNO₃) was required to cause the indicator to change color? 8.9mL
- B. What can you say about the pH before and after the last drop of HNO₃ was added?

The pH was above 8.2 when it was pink before the last drop, after the last drop it is less than 8.2

- Explore: Click Reset and change the indicator to Bromthymol blue. Add exactly 8.8 mL of HNO₃ to the flask.
 - A. What does the color of the indicator tell you about the current pH of the flask?

Since it is blue, it tells me that the pH is above 7.6

B. Add one more drop of HNO₃. What does the color tell you about the pH now?

Since it is yellow, I know that the pH is below 6.0

Activity B:	Get the Gizmo ready:	A



Titration Gizmo Answers

SB Merriam

Titration Gizmo Answers:

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