

## Atomic Emission Spectra Lab

Directions: From the lab video, record the color of flame produced by each of the metals. Use the chart provided to find the approximate wavelength of the light of each flame using the chart.

1. Before beginning, record the atomic number of the metal in each salt listed below (the metal is the first element listed).
2. Use the equation  $c = \lambda \cdot \nu$  (speed of light = frequency x wavelength) to calculate the frequency. Be sure to convert the wavelength of meters before you put it into the equation. ( $1 \text{ m} = 1 \times 10^9 \text{ nm}$ )
3. Use the frequency in the equation  $E = h \cdot \nu$  ( $h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$ ) to calculate the energy of the light.

\*\*\*\* You must show your work on all calculations to receive credit! Take a picture of your work and upload it to Canvas with your Lab. You can submit more than one document. \*\*\*

| Color        | Wavelength (nanometers) |
|--------------|-------------------------|
| Red Light    | ~650 nm                 |
| Orange Light | ~590 nm                 |
| Yellow Light | ~570 nm                 |
| Green Light  | ~510 nm                 |
| Blue Light   | ~475 nm                 |
| Indigo Light | ~445 nm                 |
| Violet Light | ~400 nm                 |



| Metal Ion Tested  | Atomic Number | Color of Flame | Wavelength (m)                  | Frequency (Hz)                  | Energy (J)             |
|-------------------|---------------|----------------|---------------------------------|---------------------------------|------------------------|
| KCl               | 19            | violet         | $4.00 \times 10^{-7} \text{ m}$ | $7.5 \times 10^{14} \text{ Hz}$ | $4.97 \times 10^{-19}$ |
| CuCl <sub>2</sub> | 29            | green          | $5.10 \times 10^{-7} \text{ m}$ | $5.9 \times 10^{14} \text{ Hz}$ | $3.91 \times 10^{-19}$ |
| CaCl <sub>2</sub> | 20            | orange         | $5.90 \times 10^{-7} \text{ m}$ | $5.1 \times 10^{14} \text{ Hz}$ | $3.38 \times 10^{-19}$ |
| NaCl              | 11            | yellow         | $5.70 \times 10^{-7} \text{ m}$ | $5.3 \times 10^{14} \text{ Hz}$ | $3.51 \times 10^{-19}$ |
| SrCl <sub>2</sub> | 38            | red            | $6.50 \times 10^{-7} \text{ m}$ | $4.6 \times 10^{14} \text{ Hz}$ | $3.05 \times 10^{-19}$ |
| LiCl              | 3             | red            | $6.50 \times 10^{-7} \text{ m}$ | $4.6 \times 10^{14} \text{ Hz}$ | $3.05 \times 10^{-19}$ |

Part 2- Describe the incandescent and fluorescent lights and how they look different through the spectrum filter.

# Virtual Chem Lab 5 Atomic Emission Spectra

**N Colangelo**



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