Name: __KEY______ Section: _____

Chapter 7 Practice Worksheet: Covalent Bonds and Molecular Structure

1) How are ionic bonds and covalent bonds different?

Ionic bonds result from the transfer of electrons from one atom to another, Covalent bonds result from two atoms sharing electrons.

Describe the relationship between the length of a bond and the strength of that bond.

Strength of a bond increases as the bond gets shorter (inverse relationship)

Identify the type(s) of bond(s) found in the following molecules:

a. CCl₄ covalent

b. Li₂O ionic

c. NF₃ covalent_____

d. CaSO₄ jonic and covalent

e. SO₂ covalent

f. Mg(OH)₂ ionic and covalent

4) Define electronegativity.

Electronegativity is the ability for an atom in a molecule to attract electrons to itself.

 Use electronegativity values to place the following elements in increasing order: F, N, H, P, Si, C, O

 $Si < C \approx P = H < N < O < F$

6) Determine if the bond between atoms in each example below is nonpolar covalent, polar covalent, or ionic.

a. H₂ npc

e. NF __pc___

b. PCI __pe____

f. MgO ionic

e. F₂ npe____

g. CH __npe__

d NaBr ionic

h. HCl pc

<u>Using Science Skills Relating Concepts Chapter 7</u> <u>Answers</u>

Lisa A. Baglione

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- 1. Understanding the eBook Using Science Skills Relating Concepts Chapter 7 Answers
 - The Rise of Digital Reading Using Science Skills Relating Concepts Chapter 7 Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Using Science Skills Relating Concepts Chapter 7 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Using Science Skills Relating Concepts Chapter 7 Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Using Science Skills Relating Concepts Chapter 7 Answers
 - Personalized Recommendations
 - Using Science Skills Relating Concepts Chapter 7 Answers User Reviews and Ratings
 - Using Science Skills Relating Concepts Chapter 7 Answers and Bestseller Lists

- 5. Accessing Using Science Skills Relating Concepts Chapter 7 Answers Free and Paid eBooks
 - Using Science Skills Relating Concepts Chapter 7 Answers Public Domain eBooks
 - Using Science Skills Relating Concepts Chapter 7 Answers eBook Subscription Services
 - Using Science Skills Relating Concepts Chapter 7 Answers Budget-Friendly Options
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 - o ePub, PDF, MOBI, and More
 - Using Science Skills Relating Concepts Chapter 7 Answers Compatibility with Devices
 - Using Science Skills Relating Concepts Chapter 7 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Using Science Skills Relating Concepts Chapter 7 Answers
 - Highlighting and Note-Taking Using Science Skills Relating Concepts Chapter 7 Answers
 - Interactive Elements Using Science Skills Relating Concepts Chapter 7 Answers
- 8. Staying Engaged with Using Science Skills Relating Concepts Chapter 7 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Using Science Skills Relating Concepts Chapter 7 Answers
- 9. Balancing eBooks and Physical Books Using Science Skills Relating Concepts Chapter 7 Answers
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Using Science Skills Relating Concepts Chapter 7 Answers
- 10. Overcoming Reading Challenges
 - o Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Using Science Skills Relating Concepts Chapter 7 Answers
 - Setting Reading Goals Using Science Skills Relating Concepts Chapter 7 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Using Science Skills Relating Concepts Chapter 7 Answers
 - Fact-Checking eBook Content of Using Science Skills Relating Concepts Chapter 7 Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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