

Understanding the Mole

Bean Lab (Relative Mass)

Introduction:

The relative mass of an object is how many times more massive the object is than a standard object. The atomic masses of atoms are all relative masses. They can be considered relative to any particular element. Historically, both oxygen and carbon have served as the reference standard. For our purposes we can also consider atomic masses relative to the least massive element--hydrogen, with an atomic mass of approx. one. Fluorine, with a relative mass of 19, is 19 times more massive than hydrogen, etc. In this laboratory exercise you deal with the relative masses of beans. Then you will be asked to draw a parallel to the atomic masses of elements.

Target:

Students are able to determine the relative mass of different small particles by comparison with an arbitrarily chosen mass. This will develop an understanding of the mole concept and molar masses of elements through an analogy with a model system.

Pre-lab assignment:

1. Read the entire lab
2. Print out the lab (data section to the end of the lab).
3. Below the target, answer the following questions regarding the lab:
 - a. Explain the process you will follow to figure out the mass of one bean.
 - b. Why should you calculate the mass of one bean rather than measuring one bean?
 - c. What should you record before you add the beans to the beaker?
 - d. In your own words, write down what is meant by "relative mass."

Procedure:

Part I. Be sure to use the same balance for the entire activity.

1. In a beaker, count out exactly 100 beans of one type. Discard any beans that differ greatly from an average bean. If you fail to do this, your results will not be accurate.
2. Put a beaker on the balance. Find the mass of the beaker. This you will need to subtract later. Add the 100 beans to the beaker; record their mass to the **nearest 0.01 g**.
3. Each team member is responsible for one type of bean provided.

Calculations:

1. Calculate (do not weigh) the mass of one bean of each type. Record the value in the data table. [Note: "Calculate" means to take the total mass of 100 beans and divide by 100 rather than weighing one particular bean.]
2. Determine the relative mass of each type of bean by comparison to the lightest type of bean. Record these values in the data table.

Relative	Mass	=	$\frac{\text{Avg. Mass of Bean}}{\text{Avg. Mass of Lightest Bean}}$
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3. Calculate the number of beans in one relative mass of each bean. Do this by dividing the relative mass by the average mass of one bean. Number of beans is = (one Relative mass/Average mass of one bean)

Understanding The Mole Bean Lab Answers

RJ Shavelson

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Understanding The Mole Bean Lab Answers:

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Comfort ,1914 **Government Reports Announcements & Index** ,1984-02 **National Gardening** ,1990

Cambridge Scientific Biochemistry Abstracts ,1992 Statement by the Secretary Smithsonian Institution,1957

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The Enigmatic Realm of **Understanding The Mole Bean Lab Answers**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing lacking extraordinary. Within the captivating pages of **Understanding The Mole Bean Lab Answers** a literary masterpiece penned by a renowned author, readers attempt a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting effect on the hearts and minds of those that partake in its reading experience.

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