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Yi-Tong Ma



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 $3\text{KClO}_3 \rightarrow 3\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $4\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $5\text{Pb} + 2\text{OH}^- + 2\text{HCl} \rightarrow 2\text{H}_2\text{O} + 1\text{PbCl}_2 + 2\text{Al} + 3\text{K}_2\text{SO}_4 + 6\text{KBr} + 1\text{Al}_2\text{SO}_4 + 3\text{H}_2\text{O}$
 $8\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 2\text{CO}_2 + 4\text{CH}_4$

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web oct 2 2019 let's start by balancing the oxygen atoms to do this make the oxygen atoms as six on either side of the chemical equation $2\text{Fe}_2\text{O}_3 \rightarrow 4\text{Fe} + 3\text{CO}_2$ on towards the next step now step 3 once you are done with this check out if the number of atoms for the other elements is also equal on both the sides

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web 1 78 348 stoichiometry helps us determine how much substance is needed or is present things that can be measured are reactants and products mass molecular weight chemical equations formulas stoichiometric coefficient the stoichiometric coefficient or stoichiometric number is the number of molecules that participate in the reaction

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web expert answer transcribed image text balancing equations and simple stoichiometry balance the following equations
 $1\text{ N}_2 + \text{F}_2 \rightarrow 2\text{NF}_3$
 $2\text{C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{GaBr} + \text{NaBr} \rightarrow \text{SNO}_4 + \text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O}$
 $5\text{NF}_3 + \text{SNF}_2 \rightarrow \text{N}_2\text{O}_3$ using the equation from problem 2 above answer the following

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web balancing equations and simple stoichiometry key balance the following equations
 $1\text{ N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$
 $2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$
 $3\text{HBr} + 1\text{KHCO}_3 \rightarrow 1\text{H}_2\text{O} + 1\text{KBr} + \text{CO}_2$
 $2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow 1\text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 6\text{NaBr} + 3\text{SNO}_2 + 3\text{NF}_3 + 3\text{SNF}_2 + 1\text{N}_2\text{O}_3$ using the following equation
 $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $\text{Ce} + \text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu} + \text{Zn} + \text{CuCl}_x + 2\text{ZnCl}_x + 2\text{Cu}$ how many moles of Ce ZnCl_2 $\text{ZnCl}_x + 2\text{Cu}$ will be produced from 23.0 g of Ce Zn Zn

assuming CeCl_2 CuCl_2 is available in excess

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 $\text{Br}_2 + \text{C}_6\text{H}_{10} \rightarrow \text{C}_6\text{H}_8 + \text{H}_2\text{Br}_2$
 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{CO}_3 + \text{H}_2\text{SO}_4 + \text{H}_2\text{O}$
br using the following equation
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$

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 $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
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 $\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{GaBr}_3 + \text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{O}_3 + \text{SO}_3 + \text{NaBr}$
 SnF_2

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web may 28 2020 write the balanced equation for the reaction and calculate how much SiO_2 is required to produce 3.00 kg of SiC
 $\text{SiO}_2 + \text{C} \rightarrow \text{SiC} + \text{CO}$

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 $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
 $\text{C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
 $\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{GaBr}_3 + \text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{O}_3 + \text{SO}_3 + \text{NaBr}$
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same thing goes with the O_3 i am so lost

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web google classroom balance the following chemical equation $\text{text Mg} + \text{OH}^- \rightarrow \text{Mg}(\text{OH})_2$ $\text{text HCl} \rightarrow \text{HCl}$ text MgCl_2 text H_2 text O_2 note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem 7 4 1 x x y y theta θ pi π 8 5 2 0 9 6

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web these consonants are voiced h w n m r j ŋ l an unvoiced consonant means that there is no vibration or voice coming from the voicebox when the sound is pronounced examples of unvoiced consonant sounds are s p and t

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