Balancing Equations Worksheet - Solutions

- 1 H₂PO₄ + 3 KOH → 1 K₂PO₄ + 3 H₂O
- 2) _6_ K + _1_ B₂O₃ → _3_ K₂O + _2_ B
- 3) _1_ HCl + _1_ NaOH → _1_ NaCl + _1_ H₂O
- 4) _10_ Na + _2_ NaNO₃ \rightarrow _6_ Na₂O + _1_ N₂
- 5) 4 C + 1 S₈ → 4 CS₂
- 6) _4_ Na + _1_ O₂ → _2_ Na₂O
- 7) $2 N_2 + 5 O_2 \rightarrow 2 N_2O_5$
- 8) $_{2}$ $_{H_{3}PO_{4}}$ + $_{3}$ $_{Mg(OH)_{2}}$ \rightarrow $_{1}$ $_{Mg_{3}(PO_{4})_{2}}$ + $_{6}$ $_{H_{2}O}$
- 9) _2_ NaOH + _1_ H₂CO₃ → _1_ Na₂CO₃ + _2_ H₂O
- 10) _1_ KOH + _1_ HBr → _1_ KBr + _1_ H₂O
- 11) _4 Na + _1 O₂ → _2 Na₂O
- 12) _2_ Al (OH)₃ + _3_ H₂CO₃ \rightarrow _1_ Al₂(CO₃)₃ + _6_ H₂O
- 13) _16_ Al + _3_ S₈ → _8_ Al₂S₃
- 14) _6_ Cs + _1_ N₂ → _2_ Cs₃N
- 15) _1_ Mg + _1_ Cl₂ → _1_ MgCl₂
- 16) _10_ Rb + _2_ RbNO₃ → _6_ Rb₂O + _1_ N₂
- 17) _2_ C₀H₀ + _15_ O₂ → _12_ CO₂ + _6_ H₂O
- 18) _1_ N₂ + _3_ H₂ → _2_ NH₃
- 19) $_{2}$ $_{10}$ H₂₂ + $_{31}$ $_{2}$ $_{2}$ $_{2}$ $_{20}$ $_{20}$ $_{20}$ + $_{22}$ $_{10}$ H₂O
- 20) $_{1}$ Al(OH)₃ + $_{3}$ HBr \rightarrow $_{1}$ AlBr₃ + $_{3}$ H₂O
- 21) _2_ $CH_2CH_2CH_3CH_3 + _13_ O_2 \rightarrow _8_ CO_2 + _10_ H_2O_2$
- 22) $_{1}$ $_{0}$ $_{1}$ $_{1}$ $_{2}$ $_{3}$ $_{4}$ $_{1}$ $_{2}$ $_{3}$ $_{4}$ $_{2}$ $_{3}$ $_{2}$ $_{3}$ $_{3}$ $_{4}$ $_{2}$ $_{3}$ $_{4}$ $_{2}$ $_{3}$ $_{4}$ $_{4}$ $_{4}$ $_{2}$ $_{3}$ $_{4$
- 23) _3_ Li + _1_ AlO₃ → _3_ LiO + _1_ Al
- 24) $_{2}$ $_{2}$ $_{3}$ $_{4}$ $_{6}$ $_{10}$ $_{2}$ $_{3}$ $_{4}$ $_{10}$ $_{2}$ $_{4}$ $_{10}$ $_{2}$ $_{2}$ $_{3}$ $_{4}$ $_{10}$ $_{2}$ $_{3}$ $_{4}$ $_{10}$ $_{2}$ $_{3}$ $_{4}$ $_{2}$ $_{2}$ $_{3}$ $_{4}$ $_{2}$ $_{2}$ $_{3}$ $_{3}$ $_{4}$ $_{2}$ $_{2}$ $_{3}$ $_{3}$ $_{4}$ $_{2}$ $_{2}$ $_{3}$ $_{3}$ $_{4}$ $_{2}$ $_{2}$ $_{3}$ $_{3}$ $_{3}$ $_{4}$ $_{2}$ $_{2}$ $_{3}$ $_{3}$ $_{3}$ $_{3}$ $_{4}$ $_{2}$ $_{3}$
- 25) _3_ NH₄OH + _1_ H₂PO₄ \rightarrow _1_ (NH₄)₃PO₄ + _3_ H₂O
- 26) _3_ Rb + _1_ P → _1_ Rb₃P
- 27) $_{1}$ $_{OH_{1}}$ + $_{2}$ $_{O_{2}}$ \rightarrow $_{1}$ $_{CO_{2}}$ + $_{2}$ $_{H_{2}O}$
- 28) $_{2}$ $_{A}(OH)_{3} + _{3}$ $_{A}(SO_{4} \rightarrow _{1}$ $_{A}(SO_{4})_{3} + _{6}$ $_{A}(OH)_{3} + _{4}(OH)_{4}$
- 29) _2_ Na + _1_ O₂ → _2_ NaO
- 30) $_{16}$ Rb + $_{1}$ S₀ \rightarrow _8 Rb₂S
- 31) _2_ $H_1PO_1 + _3_ Ca(OH)_2 \rightarrow _1_ Ca_2(PO_4)_2 + _6_ H_2O$
- 32) $_{1}$ NH₃ + $_{1}$ HCl \rightarrow $_{1}$ NH₄Cl
- 33) _2_ U + _2_ H₂O → _2_ UOH + _1_ H₂
- 34) $_1$ $_{Ca_3(PO_4)_2}$ + $_3$ $_{SiO_2}$ + $_5$ $_{C}$ \rightarrow $_3$ $_{CaSiO_3}$ + $_5$ $_{CO}$ + $_2$ $_{P}$
- 35) $_{4}$ NH₃ + $_{3}$ O₂ \rightarrow $_{2}$ N₂ + $_{6}$ H₂O
- 36) $_{\bf 4}$ FeS₂ + $_{\bf 11}$ O₂ \rightarrow $_{\bf 2}$ Fe₂O₃ + $_{\bf 8}$ SO₂
- 37) $_{5}$ C + $_{2}$ SO₂ \rightarrow $_{1}$ CS₂ + $_{4}$ CO

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Emma Smith

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that Ms Matlock Phillips had with the students in her first classroom taught her to consider their needs and the environments that they came from when teaching Administration and mentoring teachers were not able to help her to learn the individual need of her students She developed meaningful relationships with each and every one of her students because she was able to understand why they behaved in the ways that they did In other words Ms Matlock learned to respect the values that her students had learned about life and survival and developed teaching strategies that corresponded to their life lessons

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