

9 page workbook

Covers how covalent bonds work, naming covalent bonds, and drawing Lewis Dot structures of covalent bonds.

What are covalent bonds?

Covalent bonds are bonds in which one or more pairs of valence electrons are shared between the nuclei of both atoms.

A compound is produced when two or more different elements are combined covalently. When atoms bond, they produce molecules. A particle in which a covalent compound is made up of the same compound. Most compounds have a definite composition and ratio. H₂O is a compound. H₂O₂ is a different compound. H₂O and H₂O₂ have different properties.

Covalent

When two

have a

Naming Covalent Compounds

You need to know number prefixes to name covalent bonds:

One = mono-
Two = di-
Three = tri-

Drawing Covalent Bonds

To draw covalent bonds, we use Lewis Dot structures. For each atom in the molecule, use a different color or symbol to represent the valence electrons participating in the bond.

Example: MgO is a molecule made with one atom of magnesium and one atom of oxygen.

Magnesium has 2 valence electrons, and its Lewis dot structure is:



Oxygen has 6 valence electrons, and its Lewis dot structure is:



The Lewis Dot structure for MgO is:



Example: CO₂ is a molecule made with one atom of oxygen and two atoms of carbon.

Oxygen has 6 valence electrons, and its Lewis dot structure is:



element is 2
is to the
the second
atom of
oxygen.
the second
atom.

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