

Covalent Bonds

Covalent Bonds

are the forces that hold atoms together in pure compounds

Valence Electrons

are the s and p electrons in the highest occupied energy level. They are involved in shared or unshared electron bonds.

Four main types of bonds:

ionic, covalent, metallic, & hydrogen

Valence Electrons

have 8 electrons in their valence shells, which is why they will not form bonds under normal circumstances.

A chemical bond is an interaction between electrons that results in a lower potential energy of the system. This lower potential energy means that system more stable. A chemical bond of the system now has the same energy.

Covalent Bonds

Covalent bonds are attractive forces that are produced by the sharing of electrons.

Substances containing covalent bonds form molecules. Covalent bonds are between 2 non-metals and hydrogen.

eg: water

Polar Covalent vs. Non-Polar Covalent

Electronegativity greater than 1.7 = ionic bond is formed

Less than 1.7 = covalent & it can be polar or non-polar depending

Polar Covalent = unequal Non-Polar Covalent = equal

EXAMPLES: non-polar and non-polar compounds:

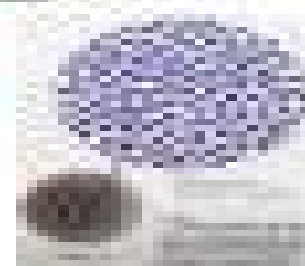
hydrocarbons like

hydrocarbons like

non-polar covalent structures are polar covalent things. Water is polar covalent and it

is non-polar covalent and that is why they don't mix.

Polar ionic compounds dissolve in polar covalent substances



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Norman H. Nachtrieb



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