CHAPTER_13: WORKSHEET (Part 1) Solubility, energetics of solution formation, concentration units, Henry's law Answer the following questions:

- a. Predict whether iodine (l₂) is more soluble in liquid ammonia (NH₂) or in carbon disulfide (CS₂).
 - b. Which of the following should you expect to be more soluble in benzene than in water CaHa, HCl. b. CSp?
 - Which of the following solutions exhibits hydrogen bonding between the solute and solvent? (Select all that apply.)
 - i Ha(g) in HaO(h
 - ii CHbOH(I) in H2O(I)
 - iii CO₂(g) in H₂O(t)
 - iv NH₂(g) in H₂O(f)
 - v. NaCl(s) in HzO(f)
- Which molecule would you expect to be more soluble in ethanol CH₂CH₂OH), CC₈ or CH₂Cl₂? Explain your choice.
- Between water and hexane (C_iH₁₄), pick an appropriate solvent to dissolve each of the following solutes:
 - a. methanol (CH_bOH).
 - b. lithium nitrate.
 - c. coconut oil.
 - d. acetic acid (CH₅COOH),
 - e. toluene (C:Ha), and
 - dimethyl ether (H₂C-O-CH₃).

State the kind of intermolecular forces that would occur between the solute and solvent in each case.

- 4. The enthalpy of solution for cesium fluoride is –36.8 kJ/mol. What can you conclude about the relative magnitudes of ΔHsolute and ΔHhydration?
- Silver nitrate has a heat of hydration of -843 kJ/mol and a heat of solution of -23 kJ/mol. What is the lattice energy of silver nitrate?
- Potassium nitrate has a lattice energy of -163.8 kcal/mol and a heat of hydration of -155.5 kcal /mol. How much potassium nitrate has to dissolve in water, at room temperature, to absorb 1.00 x 10² kJ of heat? (1 cal = 4.184 J)
- How many milliters of water (d = 0.998 g/mL) are required to dissolve 25.0 g of urea and thereby produce a 1.65 m solution of urea (CO(NH₂)₂)?
- 8. Calculate a) the molality of CH₂OH (methanol) and b) mole fraction of solvent

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Harold H. Szu, Roy F. Potter

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