



# Visual Guide To Elliot Wave

**David Baud**



## Visual Guide To Elliot Wave:

Delve into the emotional tapestry woven by Emotional Journey with in Experience **Visual Guide To Elliot Wave** . This ebook, available for download in a PDF format ( \*), is more than just words on a page; it's a journey of connection and profound emotion. Immerse yourself in narratives that tug at your heartstrings. Download now to experience the pulse of each page and let your emotions run wild.

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 $2 \text{ CO}_2 \text{ (g)} + \text{H}_2\text{O (g)} \rightarrow \text{C}_2\text{H}_2 \text{ (g)} + \frac{5}{2} \text{ O}_2 \text{ (g)}$   $\Delta H = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O (g)} + \text{H}_2 \text{ (g)} \rightarrow \frac{1}{2} \text{ O}_2 \text{ (g)}$   $\Delta H = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ (g)} + \frac{7}{2} \text{ O}_2 \text{ (g)} \rightarrow \text{CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (g)}$   $\Delta H = 283 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ (g)}$

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 $\text{C}_2\text{H}_4 \text{ (g)} + 3 \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 2 \text{ H}_2\text{O (l)}$   $\Delta H = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_6 \text{ (g)} + 7/2 \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (l)}$   $\Delta H = 1560 \text{ kJ/mole}$   
 $\text{H}_2 \text{ (g)} + 1/2 \text{ O}_2 \text{ (g)} \rightarrow \text{H}_2\text{O (l)}$   $\Delta H = 285 \text{ kJ/mole}$

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to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web thus for reaction 9 2 in which 1 mole of  $h_2o$  and 1 mole of  $oh^-$  combine to form 2 moles of  $h_2o$   $\Delta h = -55.8$  kj as shown note that the enthalpy change  $\Delta h_{products} - \Delta h_{reactants}$  is positive if heat is absorbed that is if  $\Delta h_{products}$

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