



2-2: Writing Balanced Precipitation Reactions **Lab chemistry**

In this problem, you will go into the virtual laboratory and perform a series of precipitation reactions using Ag^+ , Pb^{2+} , and Sb^{3+} . After observing the reactions, you will write the net ionic equations representing these reactions and then balance them.

1. Start *Virtual ChemLab*, select *Reactions and Stoichiometry*, and then select *Writing Balanced Precipitation Reactions* from the list of assignments. The lab will open in the Inorganic laboratory.
2. Drag a test tube from the box and place it on the metal test tube stand. You can then click on the bottle of Ag^+ ion solution (it is AgNO_3) on the shelf to add it to the test tube. Then drag the test tube to the blue test tube rack. Repeat for Pb^{2+} ($\text{Pb}(\text{NO}_3)_2$) and Sb^{3+} ($\text{Sb}(\text{OH})_3$).
3. Moving your mouse over each test tube in the test-tube rack will allow you to identify it on the TV monitor. You can also click on the label at the top of each test tube in order to label the test tubes. Place the test tube containing the Ag^+ solution in the metal test tube stand. You will be able to observe what happens in the window at the top left. Click on the Na_2CO_3 reagent bottle to add it to the test tube in the stand.

What color is the precipitate? Write out the full balanced molecular equation for the reaction.

Write the balanced ionic equation for the reaction.

Write the balanced net ionic equation for the reaction.

Virtual Lab Writing Balanced Precipitation Reactions Answers

J Dewey



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Instructors Manual to Lab Manual Ralph Petrucci, William Harwood, Geoffrey Herring, 2001

Green Chemistry

Laboratory Manual for General Chemistry Sally A. Henrie, 2015-03-18 Green chemistry involves designing novel ways to create and synthesize products and implement processes that will eliminate or greatly reduce negative environmental impacts. Providing educational laboratory materials that challenge students with the customary topics found in a general chemistry laboratory manual, this lab manual enables students to see how green chemistry principles can be applied to real world issues. Following a consistent format, each lab experiment includes objectives, prelab questions, and detailed step-by-step procedures for performing the experiments. Additional questions encourage further research about how green chemistry principles compare with traditional, more hazardous experimental methods.

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