#### Complete the following conversions. For FULL CREDIT, you must show ALL of your work!

1. Determine the number of atoms in 0.58 moles of Se.

0.58 mol Se  $\times$   $\frac{6.02 \times 10^{23} \text{ atoms Se}}{1 \text{ mol Se}} = 3.5 \times 10^{23} \text{ atoms Se}$ 

How many moles are in 6.80 x 10<sup>14</sup> formula units Ba(NO<sub>1</sub>)<sub>2</sub>?

 $6.80 \times 10^{24}$  formula units  $Ba(NO_3)_2 \times \frac{1 \ mol \ Ba(NO_3)_2}{6.02 \times 10^{23} \ formula \ units Ba(NO_3)_2} = 11.3 \ mol \ Ba(NO_3)_2$ 

Determine the number of molecules in 1.25 moles of O<sub>2</sub> gas.

 $1.25 \, mol \, O_2 \, \times \, \frac{6.02 \times 10^{23} \, molecules \, O_2}{1 \, mol \, O_2} \, = \, 7.52 \times 10^{23} \, molecules \, O_2 \, \circ$ 

4. How many moles are in 5.38 x 1024 formula units MgBr<sub>2</sub>?

 $5.38 \times 10^{24}$  formula units  $MgBr_2 \times \frac{1 \, mol \, MgBr_2}{6.02 \times 10^{23} \, formula units \, MgBr_2} = 8.94 \, mol \, MgBr_2$ 

Determine the number of formula units in 0.688 moles of AgNO<sub>3</sub>.

 $0.688 \ mol \ AgNO_{3} \times \frac{6.02 \times 10^{23} \ formula \ units \ AgNO_{3}}{1 \ mol \ AgNO_{3}} = 4.14 \times 10^{23} \ molecules \ AgNO_{3}$ 

Convert 8.46 x 10<sup>24</sup> molecules of ethane (C<sub>2</sub>H<sub>4</sub>) to moles.

 $8.46 \times 10^{24}$  molecules  $C_2H_4 \times \frac{1 \text{ mol } C_2H_6}{6.02 \times 10^{23} \text{ molecules } C_2H_6} = 14.1 \text{ mol } C_2H_6$ 

Determine the number of formula units in 1.48 moles of NaF.

1.48 mol NaF  $\times$   $\frac{6.02 \times 10^{23} formula units NaF}{1 mol NaF} = 8.91 \times 10^{23} formula units NaF$ 

Convert 8.2 x 10<sup>25</sup> molecules of water to moles.

 $8.2 \times 10^{25} \text{ molecules } H_2O \times \frac{1 \text{ mol } H_2O}{6.02 \times 10^{23} \text{ molecules } H_2O} = 140 \text{ mol } H_2O$ 

How many formula units are in 15 moles of Li₂SO<sub>4</sub>?

 $15 \ mol \ Li_2SO_4 \times \frac{6.02 \times 10^{23} \ formula \ units \ Li_2SO_4}{1 \ mol \ Li_2SO_4} = 9.0 \times 10^{24} \ formula \ units \ Li_2SO_4$ 

10. Calculate the number of moles in 7.2 x 10<sup>22</sup> atoms of sulfur.

 $7.2 \times 10^{22} \ atoms \ S \times \frac{1 \ mol \ S}{6.02 \times 10^{23} \ atoms \ S} = 0.12 \ atoms \ S$ 

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#### **Table of Contents Skills Practice 31 Moles And Formulas**

- 1. Understanding the eBook Skills Practice 31 Moles And Formulas
  - The Rise of Digital Reading Skills Practice 31 Moles And Formulas
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Skills Practice 31 Moles And Formulas
  - Exploring Different Genres
  - $\circ\,$  Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Skills Practice 31 Moles And Formulas
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Skills Practice 31 Moles And Formulas

- Personalized Recommendations
- Skills Practice 31 Moles And Formulas User Reviews and Ratings
- Skills Practice 31 Moles And Formulas and Bestseller Lists
- 5. Accessing Skills Practice 31 Moles And Formulas Free and Paid eBooks
  - Skills Practice 31 Moles And Formulas Public Domain eBooks
  - Skills Practice 31 Moles And Formulas eBook Subscription Services
  - Skills Practice 31 Moles And Formulas Budget-Friendly Options
- 6. Navigating Skills Practice 31 Moles And Formulas eBook Formats
  - o ePub, PDF, MOBI, and More
  - Skills Practice 31 Moles And Formulas Compatibility with Devices
  - Skills Practice 31 Moles And Formulas Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Skills Practice 31 Moles And Formulas
  - Highlighting and Note-Taking Skills Practice 31 Moles And Formulas
  - Interactive Elements Skills Practice 31 Moles And Formulas
- 8. Staying Engaged with Skills Practice 31 Moles And Formulas
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Skills Practice 31 Moles And Formulas
- 9. Balancing eBooks and Physical Books Skills Practice 31 Moles And Formulas
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Skills Practice 31 Moles And Formulas
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Skills Practice 31 Moles And Formulas
  - Setting Reading Goals Skills Practice 31 Moles And Formulas
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Skills Practice 31 Moles And Formulas

- Fact-Checking eBook Content of Skills Practice 31 Moles And Formulas
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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