

Gas Laws Practice

Name: _____

Date: _____

Hour: _____

IMPORTANT: whenever you use temperature, it must be in degree Kelvin (K), so remember the equation: $K = ^\circ C + 273$

1. a) convert $39\ ^\circ C$ to K. b) convert $127\ K$ to $^\circ C$.
2. A gas has an initial volume of $2.75\ L$ at a temperature of $285\ K$. If the temperature changes to $380\ K$, what is the new volume of the gas if the pressure is unchanged?
3. Gas can often be cooled by compressing it while keeping the pressure constant. If I have $45.0\ L$ of gas at room temperature ($22^\circ C$) and I compress it so that the final volume is $0.50\ L$, what is the final temperature of the gas if the pressure is constant?
4. The volume of a gas is $2.5\ L$ when the pressure is at standard pressure ($101.325\ kPa$). What is the volume of the gas if the pressure decreases to $85\ kPa$ and the temperature remains unchanged?
5. A $5.0\ L$ container of gas experiences a temperature change so that the final temperature is 4 times the initial temperature. What is the size of the container after the temperature change? (Assume constant pressure.)
6. At $45^\circ C$ the volume of a certain gas is $27.5\ L$ and the pressure is $210\ kPa$. What is the volume of the gas at standard temperature ($273\ K$) and $310\ kPa$ of pressure?
7. The pressure of a sample of gas was $97.8\ kPa$ and the volume of the gas was $3.75\ L$. If the gas occupied a container with a volume of $8.00\ L$, what would the pressure in the container be?
8. Isothermal expansion refers to allowing a gas to expand while keeping the temperature constant. This is one means to simulate a vacuum. If a gas originally at $97\ kPa$ is allowed to expand from $0.25\ L$ to $182\ L$, what is the pressure of the gas?
9. A gas is initially at a pressure of $225\ kPa$ and a temperature of $245\ K$ in a container that is $4.5\ L$. If the gas is compressed to a volume of $2.1\ L$ and the temperature changes to $275\ K$, what is the new pressure?

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