

Name: _____

Date: _____

Quantum Number Practice Worksheet

1. Summarize:

The principal quantum number, n , can have the values of: _____, etc.The angular momentum quantum number, l , can have integer values from _____ to _____.The magnetic quantum number, m_l , can have integer values from _____ to _____.2. When $n = 3$, l can have values of _____.For the 3d sublevel, l has a value of _____.When $n = 4$, l can have values of _____.For the 4p sublevel, l has a value of _____.When $n = 2$, l can have values of _____.For the 2s sublevel, l has a value of _____.

3. Summarize:

orbital	s	p	d	f
value of l				

4. There are five 4d orbitals. List the quantum numbers for each orbital.

n	l	m_l

5. Rank the following orbitals in order of increasing energy: 3s, 2s, 2p, 4s, 3p, 1s, and 3d.

6. How many orbitals in an atom can have the following quantum number or designation?

a) 3p

e) 5d

b) 4p

f) 5f

c) 4p_xg) $n = 5$

d) 6d

h) 7s

7. Answer the following questions:

a) The quantum number n describes the _____ of an atomic orbital.

b) The shape of an atomic orbital is given by the quantum number _____.

c) The maximum number of orbitals that may be associated with the set of quantum numbers $n=4$ and $l=3$ is _____.d) The maximum number of orbitals that may be associated with the quantum number set $n=3$, $l=2$, and $m_l = -2$ is _____.e) When $n=5$, the possible values of l are _____.f) The maximum number of orbitals that can be assigned to the $n=4$ shell is _____.

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