

ASSIGNMENT 3 CE SOLUTION MANUAL

- P18.13** (a) The original volume of the acetone we take as precisely 100 mL. After it is finally cooled to 20.0°C, its volume is

$$V_f = V_i(1 + \beta \Delta T) = (100 \text{ mL}) \left\{ 1 + \left[1.50 \times 10^{-4} (\text{°C})^{-1} \right] (-15.0 \text{°C}) \right\} \\ = \boxed{99.8 \text{ mL}}$$

- (b) Initially, the volume of the acetone reaches the 100-mL mark on the flask, but the acetone cools and the flask warms to a temperature of 32.0 °C. Thus, the volume of the acetone decreases and the volume of the flask increases. This means the acetone will be below the 100-mL mark on the flask.

- P18.40** (a) Particle in equilibrium model

- (b) On the piston,

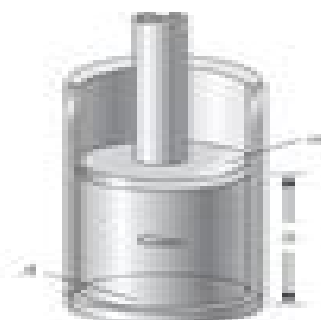
$$\sum F = F_{\text{gas}} - F_g - F_{\text{ext}} = 0 \\ \sum F = PA - mg - P_{\text{atm}}A = 0$$

- (c) In equilibrium, $P_{\text{gas}} = \frac{mg}{A} + P_{\text{atm}}$

$$\text{Therefore, } \frac{nRT}{hA} = \frac{mg}{A} + P_{\text{atm}}$$

$$\text{or } \boxed{h = \frac{nRT}{mg + P_{\text{atm}}A}}$$

where we have used $V = hA$ as the volume of the gas.



ANS. FIG. P18.40

Solution Manual Ece

Jack Philip Holman



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