

## Model 2 – Ternary Acids (Oxyacids)

| Acid                    | Name of Acid in Aqueous Solution | Cation (+)              | Polyatomic Anion (-) | Polyatomic Anion Name |
|-------------------------|----------------------------------|-------------------------|----------------------|-----------------------|
| $\text{HClO}_3$         | Chloric acid                     | $\text{H}_3\text{O}^+$  | $\text{ClO}_3^{-1}$  | Chlorate              |
| $\text{H}_2\text{SO}_3$ | Sulfurous acid                   | $2\text{H}_3\text{O}^+$ | $\text{SO}_3^{-2}$   | Sulfite               |
| $\text{H}_2\text{SO}_4$ | Sulfuric acid                    | $2\text{H}_3\text{O}^+$ | $\text{SO}_4^{-2}$   | Sulfate               |
| $\text{H}_3\text{PO}_3$ | Phosphorous acid                 | $3\text{H}_3\text{O}^+$ | $\text{PO}_3^{-3}$   | Phosphite             |
| $\text{H}_3\text{PO}_4$ | Phosphoric acid                  | $3\text{H}_3\text{O}^+$ | $\text{PO}_4^{-3}$   | Phosphate             |
| $\text{HNO}_3$          | Nitric acid                      | $\text{H}_3\text{O}^+$  | $\text{NO}_3^{-1}$   | Nitrate               |
| $\text{HNO}_2$          | Nitrous acid                     | $\text{H}_3\text{O}^+$  | $\text{NO}_2^{-1}$   | Nitrite               |
| $\text{H}_2\text{CO}_3$ | Carbonic acid                    | $2\text{H}_3\text{O}^+$ | $\text{CO}_3^{-2}$   | Carbonate             |

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

*Ternary acids contain three elements and they also contain the element oxygen.*

b. What number do you think the prefix "ter-" refers to?

*"ter" refers to three*

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming  $\text{HClO}_3$ ?

*$\text{ClO}_3^{-1}$  has the name of Chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula  $\text{HCl}$ , not  $\text{HClO}_3$ .*

10. Write a rule for naming ternary acids.

*- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.*

*- "ate" is replaced with "ic" and "ite" is replaced with "ous".*

11. Predict the formula for chlorous acid.

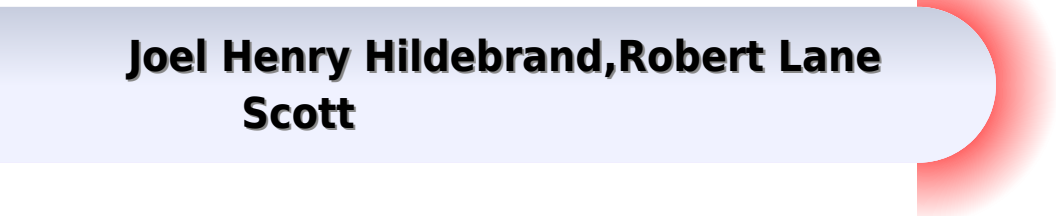


12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



# Solubility Answer Key Pogil

**Joel Henry Hildebrand, Robert Lane  
Scott**



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**Solutions and Solubilities** Michael R. J. Dack, 1975      *Principles of Solution and Solubility* Kozo Shinoda, 1991

**Solubilities of Inorganic and Organic Substances**, 1907      *Solubilities of Inorganic and Organic Substances*, 1911

**Solubilities of Inorganic and Organic Compounds** Atherton Seidell, 1919      **The Experimental Determination of**

**Solubilities** G. T. Hefter, R. P. T. Tomkins, 2003-08-01 In diesem Band werden vielfältige experimentelle Verfahren zur Bestimmung von Löslichkeiten beschrieben. Commission on Solubility Data der IUPAC beschloß die Qualität der Löslichkeitsdaten international anzugleichen und zu verbessern indem verlässliche Verfahren zur Ermittlung dieser Daten zusammengefaßt veröffentlicht werden. Alle Datentypen wurden hier berücksichtigt. Ausführliche Angaben zur Zuverlässigkeit der einzelnen Verfahren und zur Auswahl der jeweils geeigneten Methode. *Solubilities of Inorganic and Organic Substances* Atherton Seidell, 1919      *Handbook of Aqueous Solubility Data* Samuel H. Yalkowsky, Yan He, Parijat

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**Solubility Data Series**,      *Principles of Solution and Solubility* Kz Shinoda, 1978      **The Solubility of**

**Nonelectrolytes** Joel Henry Hildebrand, Robert Lane Scott, 1964      **Solubilities of inorganic and organic substances** Atherton Seidell, 1911      *Handbook of Aqueous Solubility Data* Samuel H. Yalkowsky, 2003-03-26 Over the years researchers have reported solubility data in the chemical pharmaceutical engineering and environmental literature for several thousand organic compounds. Until now this information has been scattered throughout the literature. Containing over 16 000 solubility data points for more than 4 000 organic compounds. *Handbook of Aqueous Solubility Data Series* International Union of Pure and Applied Chemistry. Commission of Equilibrium Data. Sub-Committee on Solubility Data, 1979      *Solubilities of Inorganic and Organic Compounds* Atherton Seidell, William F. Linke, 1952      **Solubilities of Inorganic and Organic**

**Compounds** Atherton Seidell, 2018-10-16 This work has been selected by scholars as being culturally important and is part of the knowledge base of civilization as we know it. This work is in the public domain in the United States of America and possibly other nations. Within the United States you may freely copy and distribute this work as no entity individual or corporate has a copyright on the body of the work. Scholars believe and we concur that this work is important enough to be preserved, reproduced and made generally available to the public. To ensure a quality reading experience this work has been proofread and republished using a format that seamlessly blends the original graphical elements with text in an easy to read typeface. We appreciate your support of the preservation process and thank you for being an important part of keeping this knowledge alive and relevant. *First Outlines of a Dictionary of Solubilities of chemical substances* Frank Humphreys STORER, 1864      **National Toxicology Program's Chemical Solubility Compendium** Lawrence H. Keith, Douglas B.

Walters,2019-11-11 This book is a compilation of experimentally determined solubility ranges of over 1 700 compounds in the National Toxicology Program s Chemical Repository Each compound s solubility was determined in a consistent manner in one to six solvents Solvents chosen were those most commonly used for toxicology studies spill cleanups and chemical synthesis or chemical reaction experiments These solvents include acetone 95% ethanol water dimethyl sulfoxide methanol and toluene Data for many of the research and industrial chemicals featured in this volume do not exist anywhere else If you are a toxicologist safety professional industrial hygienist or chemist this book is a valuable reference tool you ll find yourself using every day     Solubility curves Charles Hertzel Frantz,1933     Solubility Joel Henry Hildebrand,1924

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