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Solubility Curves

The following graph shows solubility curves. Study the graph then answer the following questions.

1. What relationship exists between solubility and temperature for most of the substances shown?

2a. What is the exception?

b. What general principle accounts for this exception?

3a. Approximately how many grams of NaNO_3 will dissolve in 100 g water at 10 °C?

b. How many grams will dissolve at 34 °C?

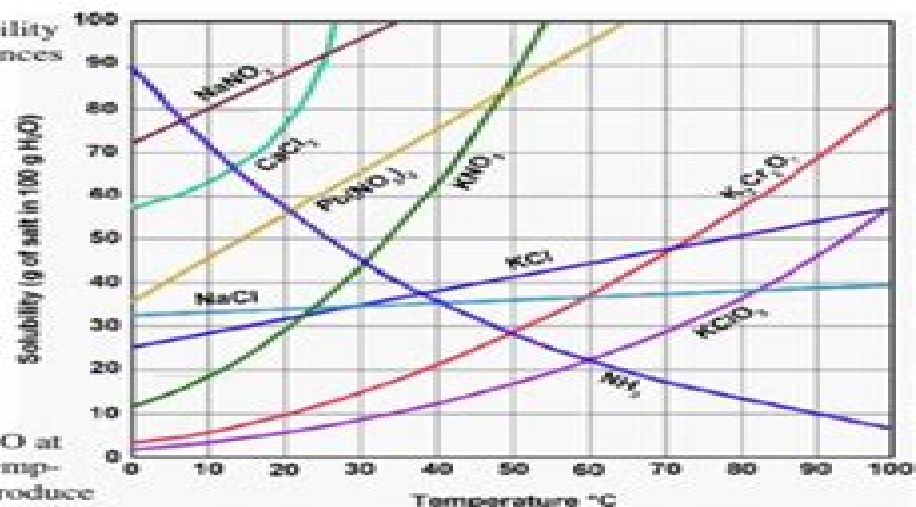
4. How many grams of KCl will dissolve in 1 liter of H_2O at 50 °C?

5. 58.0 g of $\text{K}_2\text{Cr}_2\text{O}_7$ is added to 100 g H_2O at 0 °C. With constant stirring, to what temperature must the solution be raised to produce a saturated solution with no solid $\text{K}_2\text{Cr}_2\text{O}_7$ remaining?

6. A saturated solution of KClO_3 was made with 300 g of H_2O at 34 °C. How much KClO_3 could be recovered by evaporating the solution to dryness?

7. 500 g of water are used to make a saturated solution of KCl at 15 °C. How many more grams of KCl could be dissolved if the temperature were raised to 78 °C?

8. A saturated solution of KNO_3 in 200 g of H_2O at 51 °C is cooled to 28 °C. How much KNO_3 will precipitate out of solution?



1. answer on left

2a. _____

b. answer on left

3a. _____

b. _____

4. _____

5. _____

6. _____

7. _____

8. _____

1. _____

2b. _____

Solubility Of A Salt Lab Answer Key

**James M. Postma, Julian L. Robert, J.
Leland Hollenberg**



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