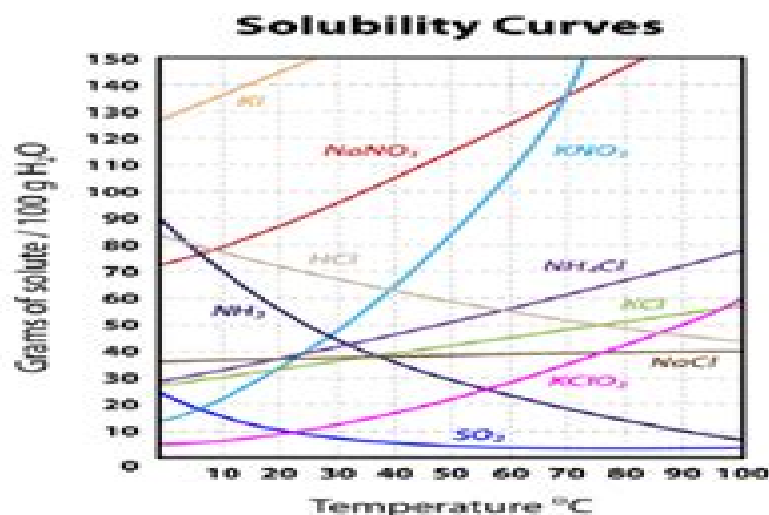


WS 6.1: Interpreting Solubility Curves

Directions: Use the Solubility Curve Table below to answer all questions. So you do not have to keep flipping your paper over you might want to put the solubility curve on your phone or computer screen when answering the questions on the back. All questions assume the amount of water is 100 ml.



- What is the solvent in this Solubility Curve?
water
- Saturation points are given per how much solvent?
100 ml or 100 g
- What is the saturation point of NaNO₃ at 10°C, 40°C, & 80°C?
@ 10°C = ~79 g
@ 40°C = ~105 g
@ 80°C = ~146 g
- How much NH₃ can you dissolve in water at 10°C, 30°C, & 90°C?
@ 10°C = ~70 g
@ 30°C = ~45 g
@ 90°C = ~10 g
- What is the maximum amount of KCl that can be dissolved in water at 20°C, 50°C, & 90°C?
@ 20°C = ~33 g
@ 50°C = ~42 g
@ 90°C = ~53 g
- If you want to make a saturated solution of NaCl how many grams would you dissolve in 100 grams of water @ 80°C?
About 40 grams
- Except for KI What substance is the most soluble in water at 20°C?
NaNO₃
- Except for KI what substance is the most soluble in water at 72°C?
KNO₃
- Look up The boiling points for SO₂, HCl, and NH₃ and write them below:
B.P. SO₂ = -10°C
B.P. HCl = -85°C
B.P. NH₃ = -33°C
- Based on the boiling points you looked up, are SO₂, HCl, & NH₃ solids, liquids, or gases at 0°C? (and anywhere else on the solubility curve)
gases

Solubility Curves Ws

Mark Steyvers

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

Solubility Curves Ws:

Journal of the American Chemical Society American Chemical Society, 1925 Proceedings of the Society are included in v 1 59 1879 1937

The Structures of Alloys of Iron W. Hume-Rothery, 2016-01-26 The Structures of Alloys of Iron An Elementary Introduction provides an elementary review of some of the factors affecting the structures of alloys of iron with examples illustrating points of interest In view of the existence of books dealing with the technology and properties of iron alloys it was decided that the present book should be concerned only with the structures and not with the properties of the alloys In connection with steels it was thought best in the space available to deal only with those based on iron carbon alloys The book begins with chapters on the metal iron the crystal structures of iron the diffusion of iron and its alloys and the intermetallic chemistry of iron Subsequent chapters deal with the structure of steels plain carbon steels the structure of alloy steels and cast irons This book is intended for those wanting to make a general survey of the subject before undertaking a detailed study The reader must understand clearly that much further work will be necessary before he can hope to understand properly even one class of the alloys of industry

Enantiomers, Racemates, and Resolutions Jean Jacques, André Collet, Samuel H. Wilen, 1994

Metamorphic Reactions and Metamorphic Facies W. S. Fyfe, 1959

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Report of Investigations, 1966

Solution and Electrolysis Sir William Cecil Dampier Dampier, 1895

The Journal of Industrial and Engineering Chemistry, 1923

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Thermodynamics and Chemistry Pierre Maurice Marie Duhem, 1903

Humidity and Moisture: Fundamentals and standards. A. Wexler and W.A. Wildhack, editors Arnold Wexler, 1965

ARS-73-1- United States. Agricultural Research Service, 1960

Fluids In The Earth's Crust W.S. Fyfe, 2012-12-02 Fluids in the Earth's Crust explores the generation and migration of fluids in the crust and their influence on the structure This book also deals with the collection and concentration of these fluids into commercially possible reservoirs or their fossil trace formed as ore bodies Chapter one of this book discusses fluid motion and geochemical and tectonic processes It then defines fluid discusses the rocks in the surface environment and provides evidence of the changes of a rock's position and the motion of fluids This book also explores the chemistry of natural fluids including the composition of ocean water pore water and deep drill fluids metamorphic fluids fluid inclusions and magmatic fluids Volatile species in minerals such as water carbon and carbon dioxide chlorine fluorine sulfur oxygen and nitrogen and other inert gases are presented in this book Other chapters in this book cover the solubility of minerals and physical chemistry of their solutions the metamorphic reactions and processes buffer systems rock deformation crustal conditions dewatering of crust and diapirism The last part of the book discusses fluids tectonics and chemical transport This book will be of great value to mining and oil geologists as well as to pure geologists

Development of Empirical Models to Predict Sucrose Crystallization in Confections Patricia A. Ryan, 2004

A Comprehensive Treatise on Inorganic and Theoretical Chemistry Joseph William Mellor, 1923

Chemical Abstracts

,1915 *Solution & Electrolysis* William Cecil Dampier Whetham,1895 Fundamentals of Engineering Donald G. Newnan,2004 Provides an in depth review of the fundamentals for the morning portion and the general afternoon portion of the FE exam Each chapter is written by an expert in the field This is the core textbook included in every FE Learning System and contains SI units Computer Techniques for Calculating Shape Factors and Channel Volumes from a Potentiometric Model for Use in Waterflood Performance Calculations Ralph V. Higgins,Dean W. Boley,Alan J. Leighton,1966 **Industrial Arts Index** ,1924

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https://movement.livewellcolorado.org/files/Resources/default.aspx/the_wounded_heart_notes.pdf

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