#### Reading a Solubility Chart

- The curve shows the # of grams of solute in a saturated solution containing 100 mL or 100 g of water at a certain temperature. (remember: 1 mL H<sub>2</sub>O = 1 g H<sub>2</sub>O)
- Any amount of solute below the line indicates an unsaturated solution at a certain temperature
- Any amount of solute above the line in which all of the solute has dissolved indicates a supersaturated solution.
- 4) If the amount of solute is above the line but has not all dissolved, the solution is saturated and the # grams of solute settled on the bottom of the container = total # g in solution - # g of a saturated solution at that temperature. (according to the curve)
- Solutes whose curves move upward w/ increased temperature are typically solids b/c the solubility of solids increases w/ increased temperature.
- 6) Solutes whose curves move downward w/ increased temperature are typically gases b/c the solubility of gases decreases with increased temperature.



#### Solubility Problems example (Many may be done by reading the solubility table)

- At 10°C, 80 g of NaNO<sub>3</sub> will dissolve in 100 mL (a saturated solution)
- To find the # grams needed to saturate a solution when the volume is NOT 100 mL use the following strategy to find answer: (e. g., 60 mL of H<sub>2</sub>O will be saturated with how many grams of NaNO<sub>3</sub>?)

Start w/ known vol. x Solubility/100mL at set temp. = amount of Solute needed to saturate

Ex. 60 mL H<sub>2</sub>O x 80 g NaNO<sub>3</sub> = 48 g NaNO<sub>3</sub> needed to saturate solution 100 mL H<sub>2</sub>O

or if the chart is in units of 100 g of H2O use the density of water conversion 1mL H2O= 1 g H2O

Ex.  $60 \text{ mL H}_2\text{O} \text{ x} \quad \frac{1 \text{ g H}_2\text{O}}{1 \text{ mL H}_2\text{O}} \text{ x} \quad \frac{80 \text{ g NaNO}_3}{100 \text{ g H}_2\text{O}} = 48 \text{ g NaNO}_3$ 

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