Table F Solubility Guidelines for Aqueous solutions

Ions That Form Soluble Compounds	Exceptions	Ions Th Insoluble C	
Group 1 ions (Li+, Na+, etc.)		carbonate (
ammonium (NH ₄ +)		chromate (
nitrate (NO ₃ ⁻)		11	
acetate (C ₂ H ₃ O ₂ ⁻ or CH ₃ COO ⁻)		phosphate (
hydrogen carbonate (HCO ₃ ⁻)		sulfide (S ² -	
chlorate (ClO ₃ ~)		hydroxide (
halides (Cl ⁺ , Br ⁺ , I ⁺)	when combined with Ag*, Pb ² *, or Hg ₂ ² *		
sulfates (SO ₄ ²⁻)	when combined with Ag*, Ca ²⁺ , Sr ²⁺ , Ba ²⁺ , or Pb ²⁺	*compounds l	

Ions That Form Insoluble Compounds*	Exceptions	
carbonate (CO ₃ ²⁻)	when combined with Group I ions or ammonium $(\mathrm{NH_4}^+)$	
chromate (CrO ₄ ²⁻)	when combined with Group 1 ions, Ca^{2+} , Mg^{2+} , or ammonium (NH_4^+)	
phosphate (PO ₄ ³⁻)	when combined with Group I tons or ammonium (NH ₄ *)	
sulfide (S ²⁻)	when combined with Group 1 ions or ammonium (NH ₄ *)	
hydroxide (OH=)	when combined with Group 1 ions, Ca ²⁺ , Ba ²⁺ , Sr ²⁺ , or ammonium (NH ₄ *)	

^{*}compounds having very low solubility in H₂O

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Huangqi Zhang

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