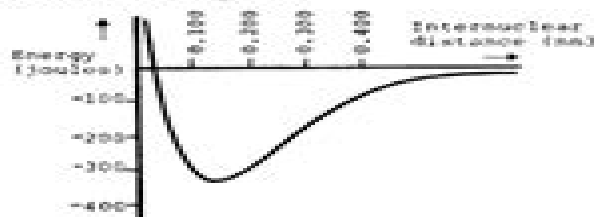


# 9 • Bonding & Molecular Structure

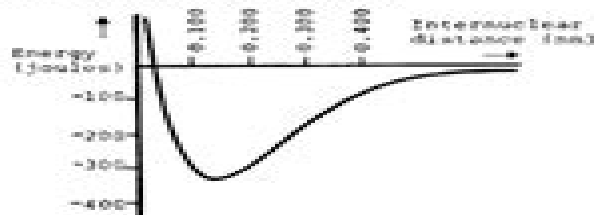
## LECTURE NOTES 1

1. A bond forms when two orbitals \_\_\_\_\_ allowing electrons to be close to two nuclei at the same time. When two H atoms bond, they overlap electrons in their \_\_\_\_\_ orbitals.
2. As two H atoms approach each other, the potential energy \_\_\_\_\_ because of the electron-proton attractions. The potential energy \_\_\_\_\_ because of electron-electron and proton-proton repulsions.
3. Consider this diagram:



Two separate atoms have an energy of \_\_\_\_ J. The equilibrium bond distance is \_\_\_\_ nm. The bond energy is \_\_\_\_ Joules which is the amount of energy you need to add to break the bond (i.e. separate the two atoms).

4. Consider this diagram again:



If two He atoms approach, their bond distance is 0.075 nm and bond energy is 100 J. Sketch the energy curve for He<sub>2</sub>. The two atoms do not get close enough for the electrons to be close to both nuclei at the same time... no bond forms.

We generalize: two \_\_\_\_\_ orbitals CAN make a bond.  
Two \_\_\_\_\_ orbitals cannot make a bond.

5. You should be able *quickly* to state the number of valence electrons any atom has based on its position on the periodic table. Write it.

|    |    |   |    |
|----|----|---|----|
| P  | Ca | F | As |
| C  | Na | K | H  |
| Si | O  | B | Mg |

6. Draw the Lewis electron-dot symbol for each of the following elements.

|    |    |   |    |
|----|----|---|----|
| P  | Ca | F | As |
| C  | Na | K | H  |
| Si | O  | B | Mg |

7. Some elements have different "ground state" Lewis symbols. Practice drawing them:

| Ground State | Bonding State | Ground State | Bonding State |
|--------------|---------------|--------------|---------------|
| C            | C             | Ca           | Ca            |
| Al           | Al            | B            | B             |

8. You should be able to recognize the **type of bond** that forms between two atoms based on the atoms involved. Classify each of the following bonds as **Ionic** or **Covalent**.

|      |       |      |       |
|------|-------|------|-------|
| SO   | _____ | KI   | _____ |
| NaCl | _____ | HCl  | _____ |
| CaBr | _____ | SiO  | _____ |
| XeF  | _____ | MgCl | _____ |
| CO   | _____ | NI   | _____ |

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**Chao Zhang**



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