

Chapter 1 - Section A - Mathcad Solutions

- 1.4 The equation that relates deg F to deg C is: $t(F) = 1.8 t(C) + 32$. Solve this equation by setting $t(F) = t(C)$.

Guess solution: $t := 0$

Given $t = 1.8t + 32$ Find(t) = -40 Ans.

- 1.5 By definition: $p = \frac{F}{A}$ $F = \text{mass} \cdot g$ Note: Pressures are in gauge pressure.
- $P := 3000 \text{ bar}$ $D := 4 \text{ mm}$ $A := \frac{\pi}{4} \cdot D^2$ $A = 12.566 \text{ mm}^2$
- $F_{\text{acc}} := P \cdot A$ $g = 9.807 \frac{\text{m}}{\text{s}^2}$ $\text{mass} := \frac{F}{g}$ $\text{mass} = 384.4 \text{ kg}$ Ans.

- 1.6 By definition: $p = \frac{F}{A}$ $F = \text{mass} \cdot g$
- $P := 3000 \text{ atm}$ $D := 0.17 \text{ m}$ $A := \frac{\pi}{4} \cdot D^2$ $A = 0.023 \text{ in}^2$
- $F := P \cdot A$ $g = 32.174 \frac{\text{ft}}{\text{sec}^2}$ $\text{mass} := \frac{F}{g}$ $\text{mass} = 1000.7 \text{ lb}_{\text{m}}$ Ans.

- 1.7 $P_{\text{abs}} = \rho \cdot g \cdot h + P_{\text{atm}}$
- $\rho := 13.535 \frac{\text{gm}}{\text{cm}^3}$ $g := 9.832 \frac{\text{m}}{\text{s}^2}$ $h := 56.38 \text{ cm}$
- $P_{\text{atm}} := 101.78 \text{ kPa}$ $P_{\text{abs}} := \rho \cdot g \cdot h + P_{\text{atm}}$ $P_{\text{abs}} = 176.808 \text{ kPa}$ Ans.

- 1.8 $\rho := 13.535 \frac{\text{gm}}{\text{cm}^3}$ $g := 32.243 \frac{\text{ft}}{\text{s}^2}$ $h := 25.62 \text{ in}$
- $P_{\text{atm}} := 29.86 \text{ in}_\text{Hg}$ $P_{\text{abs}} := \rho \cdot g \cdot h + P_{\text{atm}}$ $P_{\text{abs}} = 27.22 \text{ psia}$ Ans.

Solution Manual Chemical Engineering Kinetics

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Daneş, Valeria Petrescu, Eleonora-
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a new approach to the subject by presenting numerous applications and solved problems that illustrate the concepts introduced for varied and complex technical situations The book offers a balance between theory tools and practical applications The text aims to be a practical manual for solving engineering problems in industries where processes depend on the chemical composition and physical properties of matter The book is organized into three main topics I the molecular structure of matter II molecular models in thermodynamics and III transport phenomena and mechanisms Part I presents methods of analysis of the molecular behavior in a given system while the following parts use these methods to study the equilibrium states of a material system and to analyze the processes that can take place when the system is in a state of non equilibrium in particular the transport phenomena Molecular Physical Chemistry for Engineering Applications is designed for upper level undergraduate and graduate courses in physical chemistry for engineers applied physical chemistry transport phenomena colloidal chemistry and transport transfer processes The book will also be a valuable reference guide for engineers technicians and scientists working in industry Offers modeling techniques and tools for solving exercises and practical cases Provides solutions and conclusions so students can follow results more closely Step by step problem solving enables students to understand how to approach complex issues Kinetics of Chemical Processes Michel

Boudart,2014-05-16 Kinetics of Chemical Processes details the concepts associated with the kinetic study of the chemical processes The book is comprised of 10 chapters that present information relevant to applied research The text first covers the elementary chemical kinetics of elementary steps and then proceeds to discussing catalysis The next chapter tackles simplified kinetics of sequences at the steady state Chapter 5 deals with coupled sequences in reaction networks while Chapter 6 talks about autocatalysis and inhibition The seventh chapter describes the irreducible transport phenomena in chemical kinetics The next two chapters discuss the correlations in homogenous kinetics and heterogeneous catalysis respectively The last chapter covers the analysis of reaction networks The book will be of great use to students researchers and practitioners of scientific disciplines that deal with chemical reaction particularly chemistry and chemical engineering

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