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\* Universal gas constant  $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$ \* Avogadro constant  $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$ 

#### PART B - ESSAY

Answer two questions only. (Each question carries 150 marks.)

(a) Consider the reaction (1) given below at a temperature of 800 °C.

 $H_3(g) + I_2(g) \implies 2HI(g)$  -----(1)

Initially, 0.45 mol of HI(g) was introduced into a 1.0 dm<sup>3</sup> evacuated rigid closed container at 800 °C, and allowed to reach the equilibrium given above. It was found that 0.05 mol of  $H_2(g)$  was present at equilibrium.

- (i) Calculate the equilibrium constant K<sub>C1</sub>, for the above equilibrium at a temperature of 800 °C.
- (ii) In a separate similar evacuated container at a temperature of 800 °C, reaction (2) takes place with an equilibrium constant K<sub>C2</sub> = 1.2 × 10<sup>8</sup> mol<sup>-1</sup> dm<sup>3</sup>.

 $2H_2(g) + S_2(g) \Rightarrow 2H_2S(g) - (2)$ 

When the two containers are connected together, the following reaction (3) takes place at a temperature of 800 °C.

 $2I_2(g) + 2H_2S(g) \implies S_2(g) + 4HI(g)$  -----(3)

Calculate the equilibrium constant  $K_{C_3}$  for the reaction (3) at a temperature of 800 °C.

- (iii) A 1.0 dm³ rigid closed container at a temperature of 800 °C contains an equilibrium mixture of (3) in (ii) above with 5.00 × 10<sup>-5</sup> mol HI(g), 1.25 × 10<sup>-6</sup> mol S<sub>2</sub>(g) and 2.50 × 10<sup>-5</sup> mol H<sub>2</sub>S(g). Calculate the number of moles of I<sub>2</sub>(g) in the above mixture.
- (iv) An extra  $2.50 \times 10^{-3}$  mol  $I_2(g)$  was added to the equilibrium mixture in (iii) above at a temperature of 800 °C.
  - Calculate the reaction quotient (Q<sub>C</sub>) at the moment the extra I<sub>2</sub>(g) is added.
  - II. Explain the change in the equilibrium that takes place upon the addition of extra I2(g).
  - III. Sketch the variation in the concentrations of each constituent in the mixture, with time, upon the addition of extra I<sub>2</sub>(g).
    (60 marks)
- (b) (i) Using the data given below, calculate ΔH°, ΔS° and ΔG° for the reaction (4) at 27 °C.

 $H_2(g) + I_2(g) \longrightarrow 2HI(g)$  -----(4)

At 27 °C:  $H_2(g) + I_2(s) \rightarrow 2 HI(g)$ ;  $\Delta H^0 = 53 \text{ kJ mol}^{-1}$ ,  $\Delta S^0 = 410 \text{ J K}^{-1} \text{ mol}^{-1}$  $I_2(s) \rightarrow I_2(g)$ ;  $\Delta H^0 = 63 \text{ kJ mol}^{-1}$ ,  $\Delta S^0 = 260 \text{ J K}^{-1} \text{ mol}^{-1}$ 

(ii) Using the data given below, calculate  $\Delta H^*$ ,  $\Delta S^*$  and  $\Delta G^*$  for the reaction (5) at 27 °C.  $2 \, \text{H}_2 \text{S}(g) \longrightarrow 2 \, \text{H}_2(g) + S_2(g) ------(5)$ 

[see page ten

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**Michael Brown** 

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#### **Table of Contents Solution For Chemistry Essay 2014 2015 Question**

- 1. Understanding the eBook Solution For Chemistry Essay 2014 2015 Question
  - The Rise of Digital Reading Solution For Chemistry Essay 2014 2015 Question
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Solution For Chemistry Essay 2014 2015 Question
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Solution For Chemistry Essay 2014 2015 Question
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Solution For Chemistry Essay 2014 2015 Question
  - Personalized Recommendations
  - Solution For Chemistry Essay 2014 2015 Question User Reviews and Ratings
  - Solution For Chemistry Essay 2014 2015 Question and Bestseller Lists
- 5. Accessing Solution For Chemistry Essay 2014 2015 Question Free and Paid eBooks
  - Solution For Chemistry Essay 2014 2015 Question Public Domain eBooks
  - Solution For Chemistry Essay 2014 2015 Question eBook Subscription Services
  - Solution For Chemistry Essay 2014 2015 Question Budget-Friendly Options
- 6. Navigating Solution For Chemistry Essay 2014 2015 Question eBook Formats
  - o ePub, PDF, MOBI, and More
  - Solution For Chemistry Essay 2014 2015 Question Compatibility with Devices
  - Solution For Chemistry Essay 2014 2015 Question Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Solution For Chemistry Essay 2014 2015 Question
  - Highlighting and Note-Taking Solution For Chemistry Essay 2014 2015 Question
  - Interactive Elements Solution For Chemistry Essay 2014 2015 Question
- 8. Staying Engaged with Solution For Chemistry Essay 2014 2015 Question

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Solution For Chemistry Essay 2014 2015 Question
- 9. Balancing eBooks and Physical Books Solution For Chemistry Essay 2014 2015 Question
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Solution For Chemistry Essay 2014 2015 Question
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Solution For Chemistry Essay 2014 2015 Question
  - Setting Reading Goals Solution For Chemistry Essay 2014 2015 Question
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solution For Chemistry Essay 2014 2015 Question
  - Fact-Checking eBook Content of Solution For Chemistry Essay 2014 2015 Question
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - $\circ \ Exploring \ Educational \ eBooks$
- 14. Embracing eBook Trends
  - o Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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