

POGIL 2 Ions

How are ions made from neutral atoms?

You have learned that not all atoms of an element are the same. Variation in the number of neutrons results in different isotopes of the element. In this activity we will explore another variation that can take place—the loss and gain of electrons. The exchange of electrons between atoms is a very common way for chemical change to take place. We will see it many times throughout the year.

1. Use Diagram 1 on the next page to complete the following table.

	Metal or Nonmetal	Number of protons in the atom	Number of protons in the ion	Number of neutrons in the atom	Number of neutrons in the ion	Number of electrons in the atom	Number of electrons in the ion	Charge on the ion
Lithium	metal	3	3	4	4	3	2	1+
Magnesium	<i>Metal</i>	12	12	12	12	12	10	2+
Aluminum	<i>Metalloid</i>	13	13	14	14	13	10	3+
Fluorine	<i>Non metal</i>	9	9	10	10	9	10	1-
Oxygen	nonmetal	8	8	8	8	8	10	2-
Nitrogen	<i>Non metal</i>	7	7	7	7	7	10	3-

2. Based on the table you completed in Question 1, what distinguishes a neutral atom from an ion?

A neutral atom has the same number of protons and electrons (no charge). An ion has a charge (# of protons does not equal # of electrons).

3. Examine the isotope symbols in Diagram 1.

- a. Where is the ion charge located in the isotope symbol?

Top right corner

- b. Is a charge indicated on the neutral atoms? If yes, where is it located?

No

4. Which subatomic particle carries a positive charge?

Protons are positive

5. Which subatomic particle carries a negative charge?

Electrons are negative

6. Propose a mathematical equation to calculate the charge on an ion from the number of protons and electrons in an ion. Confirm that your equation works using two positive ion examples and two negative ion examples from Diagram 1.

The charge of an ion = # of protons - # of electrons

Solubility Packet Answers Pogil

**Joel Henry Hildebrand, Robert Lane
Scott**



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Solubility data series International Union of Pure and Applied Chemistry. Commission on Solubility Data, 1979

Handbook of Aqueous Solubility Data Samuel H. Yalkowsky, Yan He, Parijat Jain, 2016-04-19 Over the years researchers have reported solubility data in the chemical pharmaceutical engineering and environmental literature for several thousand organic compounds Until the first publication of the Handbook of Aqueous Solubility Data this information had been scattered throughout numerous sources Now newly revised the second edition of Regular and Related Solutions Joel Henry Hildebrand, 2000 **Solubility of Gases and Liquids** W. Gerrard, 2013-11-11 The solubility of gases and liquids in liquids is of great importance in large areas of operations based on chemical concepts Phenomena have appeared to be so varied that even experts have from time to time remarked on the difficulty of seeing a consistent pattern Now for the first time the essential pattern of all known gas solubility data is set out in a graphic form for all to see The continuous merging of the gas liquid systems and the liquid liquid systems is also illustrated The pattern opens the way to rational predictions The new data given for the lower alkanes and alkenes the three methylamines ammonia bromomethane and chloroethane together with my previously reported data on hydrogen sulfide dimethyl ether chloromethane and sulfur dioxide have been obtained by a bubbler manometer procedure which is fully described Not only are these data of significance in many chemical processes but they have also been vital to the development of the overall essential pattern covering all gases The book is for chemists chemical engineers biotechnologists certain physicists and teachers and students in these disciplines It is a book for all those who are concerned with the use and inculcation of the fundamental even rudimentary principles of chemistry *Developments and Applications in Solubility* Trevor M Letcher, 2007-10-31 Solubility is fundamental to most areas of chemistry and is one of the most basic of thermodynamic properties It underlies most industrial processes Bringing together the latest developments and ideas *Developments and Applications in Solubility* covers many varied and disparate topics The book is a collection of work from leading experts in their fields and covers the theory of solubility modelling and simulation industrial applications and new data and recent developments relating to solubility Of particular interest are sections on experimental calculated and predicted solubilities solubility phenomena in green quaternary mixtures involving ionic liquids molecular simulation approaches to solubility solubility impurities in cryogenic liquids and carbon dioxide in chemical processes The book is a definitive and comprehensive reference to what is new in solubility and is ideal for researcher scientists industrialists and academics **Solubilities of Inorganic and Organic Substances**, 1907

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