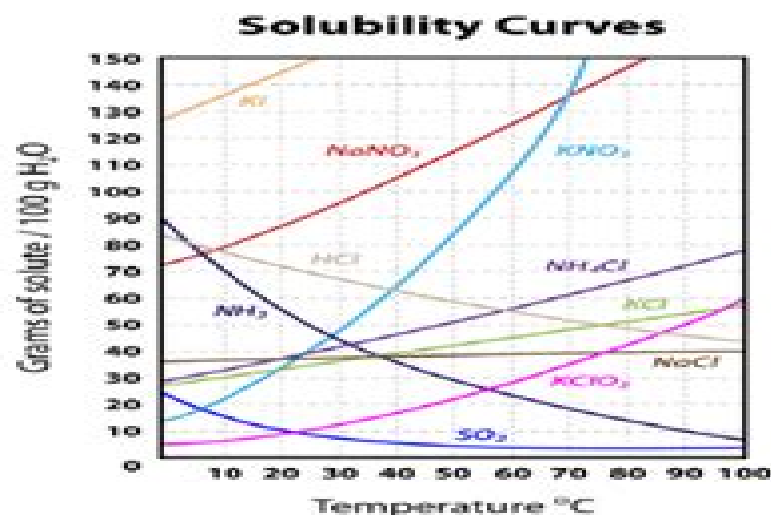


WS 6.1: Interpreting Solubility Curves

Directions: Use the Solubility Curve Table below to answer all questions. So you do not have to keep flipping your paper over you might want to put the solubility curve on your phone or computer screen when answering the questions on the back. All questions assume the amount of water is 100 ml.



- What is the solvent in this Solubility Curve?
water
- Saturation points are given per how much solvent?
100 ml or 100 g
- What is the saturation point of NaNO_3 at 10°C , 40°C , & 80°C ?
@ 10°C = ~79 g
@ 40°C = ~105 g
@ 80°C = ~146 g
- How much NH_3 can you dissolve in water at 10°C , 30°C , & 90°C ?
@ 10°C = ~70 g
@ 30°C = ~45 g
@ 90°C = ~10 g
- What is the maximum amount of KCl that can be dissolved in water at 20°C , 50°C , & 90°C ?
@ 20°C = ~33 g
@ 50°C = ~42 g
@ 90°C = ~53 g
- If you want to make a saturated solution of NaCl how many grams would you dissolve in 100 grams of water @ 80°C ?
About 40 grams
- Except for KCl What substance is the most soluble in water at 20°C ?
 NaNO_3
- Except for KCl what substance is the most soluble in water at 72°C ?
 KNO_3
- Look up The boiling points for SO_2 , HCl , and NH_3 and write them below:
B.P. SO_2 = -10°C
B.P. HCl = -85°C
B.P. NH_3 = -33°C
- Based on the boiling points you looked up, are SO_2 , HCl , & NH_3 solids, liquids, or gases at 0°C ? (and anywhere else on the solubility curve)
gases

Solubility Curves Answers

Kaplan Test Prep

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the Kaplan Test Prep text.

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Table of Contents Solubility Curves Answers

1. Understanding the eBook Solubility Curves Answers
 - The Rise of Digital Reading Solubility Curves Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Solubility Curves Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solubility Curves Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solubility Curves Answers
 - Personalized Recommendations
 - Solubility Curves Answers User Reviews and Ratings
 - Solubility Curves Answers and Bestseller Lists

5. Accessing Solubility Curves Answers Free and Paid eBooks
 - Solubility Curves Answers Public Domain eBooks
 - Solubility Curves Answers eBook Subscription Services
 - Solubility Curves Answers Budget-Friendly Options
6. Navigating Solubility Curves Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Solubility Curves Answers Compatibility with Devices
 - Solubility Curves Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solubility Curves Answers
 - Highlighting and Note-Taking Solubility Curves Answers
 - Interactive Elements Solubility Curves Answers
8. Staying Engaged with Solubility Curves Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solubility Curves Answers
9. Balancing eBooks and Physical Books Solubility Curves Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solubility Curves Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solubility Curves Answers
 - Setting Reading Goals Solubility Curves Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solubility Curves Answers
 - Fact-Checking eBook Content of Solubility Curves Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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