Answer Keys SCH4U1 Unit 5 - Thermochemistry and Rates of Reaction Summative Assessment K/U: TH: APP Please answer the following questions in the space provided. [K5] a) When solid KBr is dissolved in water, the endothermic solution gets colder. Is the chemical system exothermic or endothermic? b) When a sample of potassium nitrate is dissolved in water, the temperature of the endathermic water changes from 23.5°C to 20.3°C. Is the chemical system exothermic or [47 endothermic? c) In the reaction 2A + B → C + 3D. reactant A is found to disappear at the rate of 1 mel of B disappears at x mol 6.2 x 10⁻⁴ moVL.s. What is the rate of disappearance of B? d) The following diagram shows a collision between two BrNO molecules. Use collision theory to explain the results of the collision. Reaction did took place. New compounds formed: 2 NO and 1Br are the products. It was an effective collision. Reactants collided with proper orientation and sufficient Consider the following reaction: 2 C₂H₂₀₀ + 5 O₂₀₀ → 4 CO₂₀₀ + 2 H₂O₅₀ $\Delta H = -2599 \text{ kJ}$ a) Write a thermochemical equation with the enthalpy change, ΔH, 'inside' the equation. [C1] 2 C2 H2(g) + 50z(g) -> 4 COz(g) + 2 H2O(e) + 2599 KT b) Sketch an enthalpy diagram for this reaction. Provide appropriate labels for each axes [C3] 1 2 C2 H2/8) + 5 O2/9) Enthaloy reactants 4H=-2599KT 4002(g) + 2H2 (11) Reaction coordina

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George I. Sackheim, Dennis D. Lehman, Ronald M. Schultz

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