

Name : _____ Date : _____

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Solubility Rules Practice Worksheet

1. Classify each substance as being soluble or insoluble in water.

I. $\text{Mg}(\text{PO}_4)_2$ - _____

VI. KBr - _____

II. KOH - _____

VII. $\text{Pb}(\text{CO}_3)_2$ - _____

III. NiCl_2 - _____

VIII. PbI_2 - _____

IV. NH_4OH - _____

IX. BaSO_4 - _____

V. Hg_2SO_4 - _____

X. NiCl_2 - _____

2. Show the ions that formed the following compounds:

I. $\text{Zn}_3(\text{PO}_4)_2$

II. Al_2S_3

III. Iron (III) sulfide

IV. Ammonium cyanide

3. Form 4 water-soluble compounds by combining ions from the ions below:

Cl^-

CO_3^{2-}

PO_4^{3-}

Li^+

Sr^{2+}

4. Identify the precipitate in the following reaction. Circle the correct answer.



5. Name or give the chemical formula for each of the following compounds and state whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution.

Chemical Formula	Name	Solubility
$\text{Zn}_3(\text{PO}_4)_2$		
	Silver bromide	
KNO_3		
	Aluminum sulfide	
	Silver acetate	

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Table of Contents Solubility Practice Problems

1. Understanding the eBook Solubility Practice Problems
 - The Rise of Digital Reading Solubility Practice Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Solubility Practice Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solubility Practice Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solubility Practice Problems
 - Personalized Recommendations
 - Solubility Practice Problems User Reviews and Ratings
 - Solubility Practice Problems and Bestseller Lists
5. Accessing Solubility Practice Problems Free and Paid eBooks
 - Solubility Practice Problems Public Domain eBooks
 - Solubility Practice Problems eBook Subscription Services
 - Solubility Practice Problems Budget-Friendly Options
6. Navigating Solubility Practice Problems eBook Formats

- ePub, PDF, MOBI, and More
- Solubility Practice Problems Compatibility with Devices
- Solubility Practice Problems Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solubility Practice Problems
 - Highlighting and Note-Taking Solubility Practice Problems
 - Interactive Elements Solubility Practice Problems
- 8. Staying Engaged with Solubility Practice Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solubility Practice Problems
- 9. Balancing eBooks and Physical Books Solubility Practice Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solubility Practice Problems
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Solubility Practice Problems
 - Setting Reading Goals Solubility Practice Problems
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solubility Practice Problems
 - Fact-Checking eBook Content of Solubility Practice Problems
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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