

Solubility Rules

When salt solutions are mixed, two things can happen: 1) no reaction (all ions remain soluble), or 2) a solid forms (called a precipitate). You will examine which salts (positive + negative ions) are soluble and which are insoluble.

Safety: Wear goggles, wash your hands after lab.

Hypothesis: Look at the grid below. Looking back at your solubility rules sheet, which of these reactions do you expect to form a precipitate (solid)? _____

Procedure

1. Obtain from teacher: 1 well plate and aqueous solutions to be used in investigations.
2. Place your well plate on the lab bench.
3. Using the dropper bottles at your lab station, carefully add one drop (from a height of 1-2 cm) to each well for that ion. E.g. for Br^- a drop goes in each box under Br^- ; for Ag^+ a drop goes in each box left of Ag^+ . (be careful to note the label on each bottle). It is critical to add drops from a height of 1-2 cm. If you touch the tip of the dropper bottle to another solution, the contents of that bottle can become contaminated, ruining the results of your group and other students. Continue this procedure until all ions have been added together. At the end of this step you should have, in each box, a mixture of two drops (1 positive ion + 1 negative ion).
4. Look in the boxes: some will contain a precipitate (i.e. solid), others will be clear (indicating the salt is soluble).
5. On the grid below, place an "X" in every box where a precipitate has formed.
6. Rinse off the reaction grid in the sink. Rinse out the sink. Wipe the grid clean using a damp paper towel. Dry the grid. Return all equipment. Wipe off your lab bench. Wash your hands thoroughly.

	AgNO_3 (Ag^+)	$\text{Pb}(\text{NO}_3)_2$ (Pb^{+2})	CaCl_2 (Ca^{+2})	CuSO_4 (Cu^{+2})	$\text{Fe}(\text{NO}_3)_3$ (Fe^{+3})	K_2CrO_4 (K^+)
Na_3PO_4 (PO_4^{-3})	a	d	g	j	m	p
NaOH (OH^-)	b	e	h	k	n	q
Na_2SO_4 (SO_4^{-2})	c	f	i	l	o	r

Solubility Rules And The Mystery Solutions Lab Answers

Xiaolong Qi



Solubility Rules And The Mystery Solutions Lab Answers:

Study Guide/Selected Solutions Manual Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

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In a digitally-driven earth where screens reign great and quick transmission drowns out the subtleties of language, the profound secrets and emotional subtleties concealed within phrases frequently go unheard. However, nestled within the pages of **Solubility Rules And The Mystery Solutions Lab Answers** a captivating literary treasure pulsating with natural thoughts, lies an exceptional journey waiting to be undertaken. Penned by a talented wordsmith, this charming opus encourages viewers on an introspective journey, lightly unraveling the veiled truths and profound impact resonating within the cloth of each and every word. Within the mental depths with this poignant review, we shall embark upon a genuine exploration of the book's core subjects, dissect its interesting publishing type, and succumb to the effective resonance it evokes strong within the recesses of readers' hearts.

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