Solubility Product, Ksp - Worksheet 1

1) Write the dissociation equation and the solubility product (Ksp) expression for the dissociation of each of the following substances.

a) PbSO₄ (s) b) Mg(OH)₂ (s) c) Pb(NO₃)₂ (s)d) AlCl₃ (s)

e) Na₂S (s) f) Ca₃(PO₄)₂(s)

Write the solubility product expression for the dissociation of each of the following solids in water.

a) sodium chloride

c) silver sufide

e) magnesium sulfite

b) calcium carbonate

d) aluminum hydroxide

f) barium phosphate

The solubility of NiS is 1.0 x 10⁻¹² mol/L. Calculate the value of K_{sp} for this compound.

4) The solubility of calcium hydroxide (Ca(OH)₂) is 0.02 mol/L. Calculate the solubility product constant for this substance.

5) The solubility of silver sulfide (Ag₂S) is 1.3 x 10⁻⁶ mol/L. Calculate the K_{sp} value.

6) Experiments show that 0.0059 g of SrCO₃ will dissolve in 1.0 L of water. What is the K_{8P} for strontium carbonate?

7) The solubility of cobalt (II) sulfide is 3.8 x 10⁻³ g/L. Calculate the value of K_{sp}.

Solubility Product, K_{sp} - Worksheet 2

1. Calculate the solubility (in mol/L) for each of the following compounds:

a. FeS $K_{sp} = 3.7 \times 10^{-19}$ b. Mg(OH)₂ $K_{So} = 9.0 \times 10^{-12}$

b. Mg(OH)₂ Ksp = 9.0 x 10⁻¹²
 c. Ag₂S Ksp = 1.6 x 10⁻⁴⁹

d. AgCN K_{sp} = 2.0 x 10⁻¹²

- 2. What is the solubility of silver iodide in grams per litre? $(K_{sp} = 8.3 \times 10^{-17})$
- The K_{sp} value for BaSO₄ is 1.5 x 10⁻⁹. Calculate the number of moles of this salt that will dissolve in 1000 mL of water.
- 4. Calculate the mass of CuS that will dissolve in 1.0 litre of water. $K_{sp} = 4.0 \times 10^{-38}$
- Calculate the mass of barium carbonate that will dissolve in 100 mL of water.
 The K_{8P} of BaCO₃ is 8.1 x 10⁻⁹.

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Sherry Lewin

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