A 1.000 g sample of decane,  $C_{10}H_{22}$ , is ignited inside a bomb calorimeter. The temperature rises from 20.0°C to 78.8°C. Given that the heat capacity of the calorimeter is 810.1 J/°C, determine the heat (q) for the combustion of decane in units of kJ/mol of decane.

$$(1)(800.4)(58.8)$$
  
 $(9=47633.9)$ 

Calculate the final temperature after 1485 J of energy is added to 16.7 g of water at 23.4°C.

$$q = mCAT$$
 $1485 = (16.7)(4.184)(T_{+}-23.4)$ 
 $1485 = (69.8728)(T_{+}) - 1635.074$ 
 $3120.024 = 69.8728T_{+}$ 

How much heat is absorbed by a copper penny with a mass of 3.10 g when its temperature rises from -8.0 °C to 37 °C?

$$T_{i} = 37$$
 $T_{i} = -8.0$ 
 $37 + 8 = 45^{\circ}C$ 
 $(3.19)(45)(0.385) = 53.71 T$ 

An unknown metal has a mass of 19.7 g. When 8.20 x 10<sup>2</sup> J of heat are added to the sample, the sample temperature changes by 55.5 °C. Calculate the specific heat capacity of the unknown metal.

mc
$$\Delta T$$
  
m=19.7 (19.7)(c)(55.5)=820  
C=  
 $\Delta T = 55.5$   
 $q = 820 T$ 

# **Specific Heat Calculations With Answers**

E. N. Ramsden

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