

Specific Heat Capacity Lab Practice

Specific Heat Capacity_{water} = $4.184 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}$ $q = m \cdot c \cdot \Delta T$ or $c = \frac{q}{m \cdot \Delta T}$

A 25.00 g piece of metal at 110.0°C is placed in a calorimeter containing 15.00 g of water at 25.00°C . The final temperature of the system is 40.70°C . Determine the specific heat capacity of the metal.

Heat gained by water = $q = m \cdot c \cdot \Delta T$ or $c = \frac{q}{m \cdot \Delta T}$ or $q = m \cdot c \cdot \Delta T$
 is equal to
 $m = 15.00 \text{ g}$ $\Delta T = 15.70^\circ\text{C}$

Heat lost by metal = $q = m \cdot c \cdot \Delta T$ $m = 25.00 \text{ g}$ $\Delta T = 70.30^\circ\text{C}$
 $\frac{q}{m \cdot \Delta T} = \frac{4.184 \text{ J}}{\text{g} \cdot ^\circ\text{C}}$ $\frac{q}{25.00 \text{ g} \cdot 70.30^\circ\text{C}} = \frac{4.184 \text{ J}}{15.00 \text{ g} \cdot 15.70^\circ\text{C}}$
 $\boxed{c = 0.2583 \text{ J/g}^\circ\text{C}}$ Answer

Specific Heat Capacity_{water} = $4.184 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}$

A 40.0 Gg piece of metal at 100.0°C is placed in a calorimeter containing 30.00 g of water at 20.00°C . The final temperature of the system is 25.5°C . Determine the specific heat capacity of the metal.

Heat gained by H_2O = heat lost by metal
 Heat gained by H_2O = $q = m \cdot c \cdot \Delta T$ or $c = \frac{q}{m \cdot \Delta T}$ or $q = m \cdot c \cdot \Delta T$
 $m = 30.00 \text{ g}$ $\Delta T = 5.5^\circ\text{C}$

$\frac{q}{m \cdot \Delta T} = \frac{4.184 \text{ J}}{\text{g} \cdot ^\circ\text{C}}$ $\frac{q}{40.00 \text{ g} \cdot 80.0^\circ\text{C}} = \frac{4.184 \text{ J}}{30.00 \text{ g} \cdot 5.5^\circ\text{C}}$ $\frac{100.0^\circ\text{C}}{20.0^\circ\text{C}}$
 $\boxed{c = 0.2583 \text{ J/g}^\circ\text{C}}$ Answer

Specific Heat Capacity Lab Manual Answers

**Neena Sinha, R Rangarajan, R P
Manchanda, R K Gupta, Rajesh Kumar**



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LK-Science-HB-09-R R Rangarajan, Neena Sinha, Rajesh Kumar, LK Science HB 09 R

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