

**University of Mumbai**  
**Revised syllabus S.Y.B.Sc in Chemistry**  
**(w.e.f. the academic year 2009 – 10)**

**Paper I**

Total Lectures 90

Physical and Industrial Analytical Chemistry  
(Numerical problems are expected on all units except Unit 6)

**Term – I**  
**Unit I**

**Chemical Thermodynamics**

1. Gibbs free energy and Helmholtz's free energy, variation of Gibbs free energy with temperature and pressure, Gibbs – Helmholtz equation.
2. Physical equilibria involving pure substances, Clapeyron's equation, and variation of vapour pressure with temperature, Clausius – Clapeyron equation and its application. Partial molal properties, partial molal volume and chemical potential, Gibbs – Duhem equation.
3. Variation of chemical potential with pressure and temperature, fugacity, activity and their relationship with chemical potential, activity and activity coefficient.
4. Thermodynamic derivation of Law of mass action,  $K_p$ ,  $K_c$  and their interrelation, van't Hoff's reaction isotherm and reaction isochore.

**Unit – II**

**Solution of electrolytes:**

1. Electronic and electrolytic conductors, conductance, specific conductance, equivalent conductance, molar conductance, determination of molar conductance, variation of molar conductance with concentration for strong and weak electrolytes, concept of limiting molar conductance.
2. Debye – Huckel theory of conductance of strong electrolytes, ionic atmosphere relaxation effect, electrophoretic effect.
3. Kohlrausch's law of independent migration of ions.  
Applications of Kohlrausch's law
  - a. Determination of limiting molar conductance of a weak electrolyte,
  - b. Determination of dissociation constant of a weak acid
  - c. Determination of solubility product of a sparingly soluble salt
2. Migration of ions, transport number, dependence of transport number on the velocity of the ion, Hittorf's rule for unattackable electrodes illustrate with ions having equal and the unequal speeds, moving boundary method for the determination of the transport number, factors affecting the transport number of an ion.

# Syllabus Of Sybsc Chemistry

**Rosina Ehmman**



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