

Practice Problems

1. Select the term in Column B that best matches the phrase in Column A.

Column A

- a. Cathode ray
- b. Discovered in 1932
- c. Caused large deflections of alpha particles in Rutherford's experiment
- d. Has a charge of $1-$
- e. Has no charge
- f. Contains nearly all of an atom's mass
- g. In an atom, the number of these particles is equal to the number of protons.
- h. Identified by Thomson
- i. Site of an atom's positive charge
- j. Has a positive charge and relative mass of 1
- k. The center of an atom
- l. Symbolized by n^0

Column B

- 1. electron
- 2. proton
- 3. neutron
- 4. nucleus

4.3 How Atoms Differ

It is important to be able to obtain useful information from the periodic table. Starting with the simplest element, hydrogen, the periodic table represents each element in a box.

► **Information about an element** The name of the element and its chemical symbol are given in each box. The integer in the box represents the element's **atomic number**, the number of protons in a nucleus of an atom of that element. As you scan the table, you can see that the elements are arranged in order of increasing atomic number and that the sequence of atomic numbers has no missing values. This continuous series is an indication of the fact that the atomic number determines the identity of an element.

Solving Problems A Chemistry Handbook Answers Key

Chapter 8

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