

Chemistry

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(Chapter - 3) (Chemical Kinetics) (Exercises)
(Class 12)

Question 3.1:

From the rate expression for the following reactions, determine their order of reaction and the dimensions of the rate constants.



Answer 3.1:



Therefore, order of the reaction = 2

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{NO}]^2} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{(\text{mol L}^{-1})^2} = \frac{\text{mol L}^{-1}\text{s}^{-1}}{\text{mol}^2\text{L}^{-2}} = \text{L mol}^{-1}\text{s}^{-1}$$



Therefore, order of the reaction = 2

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{H}_2\text{O}_2][\text{I}^-]} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{(\text{mol L}^{-1})(\text{mol L}^{-1})} = \text{L mol}^{-1}\text{s}^{-1}$$



Therefore, order of reaction = $3/2$

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{CH}_3\text{CHO}]^{3/2}} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{(\text{mol L}^{-1})^{3/2}} = \frac{\text{mol L}^{-1}\text{s}^{-1}}{\text{mol}^{3/2}\text{L}^{3/2}} = \text{L}^{3/2}\text{mol}^{1/2}\text{s}^{-1}$$



Therefore, order of the reaction = 1

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{C}_2\text{H}_5\text{Cl}]} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{\text{mol L}^{-1}} = \text{s}^{-1}$$

Question 3.2:

For the reaction:



The rate = $k[\text{A}][\text{B}]^2$ with $k = 2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1}$. Calculate the initial rate of reaction when $[\text{A}] = 0.1 \text{ mol L}^{-1}$, $[\text{B}] = 0.2 \text{ mol L}^{-1}$. Calculate the rate of reaction after $[\text{A}]$ is reduced to 0.06 mol L^{-1} .

Answer 3.2:

The initial rate of reaction is

$$\text{Rate} = k[\text{A}][\text{B}]^2 \Rightarrow (2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1})(0.1 \text{ mol L}^{-1})(0.2 \text{ mol L}^{-1})^2 = 8.0 \times 10^{-8} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1}$$

When $[\text{A}]$ is reduced from 0.1 mol L^{-1} to 0.06 mol L^{-1} ,

$$\text{The concentration of A reacted} = (0.1 - 0.06) \text{ mol L}^{-1} = 0.04 \text{ mol L}^{-1}$$

$$\text{Therefore, concentration of B reacted} = \frac{1}{2} \times 0.04 \text{ mol L}^{-1}$$

$$\text{Then concentration of B available, } [\text{B}] = (0.2 - 0.02) \text{ mol L}^{-1} = 0.18 \text{ mol L}^{-1}$$

After $[\text{A}]$ is reduced to 0.06 mol L^{-1} , the rate of reaction is given by,

$$\text{Rate} = k[\text{A}][\text{B}]^2 \Rightarrow (2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1})(0.06 \text{ mol L}^{-1})(0.18 \text{ mol L}^{-1})^2 = 3.89 \text{ mol L}^{-1} \text{ s}^{-1}$$

Question 3.3:

The decomposition of NH_3 on platinum is zero order reaction. What are the rates of production of N_2 and H_2 if $k = 2.5 \times 10^{-4} \text{ mol}^{-1} \text{ L s}^{-1}$?

Answer 3.3:

The decomposition of NH_3 on platinum surface is represented by the following equation.



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**Frank Moore Colby, Allen Leon
Churchill, Herbert Treadwell
Wade, Frank H. Vizetelly**



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The Science of Solar System Ices Murthy S. Gudipati, Julie Castillo-Rogez, 2012-07-28 The role of laboratory research and simulations in advancing our understanding of solar system ices including satellites KBOs comets and giant planets is becoming increasingly important Understanding ice surface radiation processing particle and radiation penetration depths surface and subsurface chemistry morphology phases density conductivity etc are only a few examples of the inventory of issues that are being addressed by Earth based laboratory research As a response to the growing need for cross disciplinary dialog and communication in the Planetary Ices science community this book aims to achieve direct dialog and foster focused collaborations among the observational modeling and laboratory research communities

Physics Briefs, 1988

The Brush and the Pen Dario Gamboni, 2011 French symbolist artist Odilon Redon 1840-1916 seemed to thrive at the intersection of literature and art Known as the painter writer he drew on the works of Poe Baudelaire Flaubert and Mallarmé for his subject matter And yet he concluded that visual art has nothing to do with literature Examining this apparent contradiction *The Brush and the Pen* transforms the way we understand Redon's career and brings to life the interaction between writers and artists in fin de siècle Paris Dario Gamboni tracks Redon's evolution from collaboration with the writers of symbolism and decadence to a defense of the autonomy of the visual arts He argues that Redon's conversion was the symptom of a mounting crisis in the relationship between artists and writers provoked at the turn of the century by the growing power of art criticism that foreshadowed the modernist separation of the arts into intractable fields In addition to being a distinguished study of this provocative artist *The Brush and the Pen* offers a critical reappraisal of the interaction of art writing criticism and government institutions in late nineteenth century France

Constitutive Equations American Society of Mechanical Engineers. Winter Annual Meeting, 1984

Mining Journal, 1924

Federal Register, 2007-07-27

Sea Ice David N. Thomas, 2017-03-06 Over the past 20 years the study of the frozen Arctic and Southern Oceans and sub arctic seas has progressed at a remarkable pace This third edition of *Sea Ice*

gives insight into the very latest understanding of the how sea ice is formed how we measure and model its extent the biology that lives within and associated with sea ice and the effect of climate change on its distribution How sea ice influences the oceanography of underlying waters and the influences that sea ice has on humans living in Arctic regions are also discussed Featuring twelve new chapters this edition follows two previous editions 2001 and 2010 and the need for this latest update exhibits just how rapidly the science of sea ice is developing The 27 chapters are written by a team of more than 50 of the worlds leading experts in their fields These combine to make the book the most comprehensive introduction to the physics chemistry biology and geology of sea ice that there is This third edition of Sea Ice will be a key resource for all policy makers researchers and students who work with the frozen oceans and seas

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